MOI HIGH SCHOOL KABARAK

TRAIL 1 2024

Kenya Certificate of Secondary Education

MARKING SCHEME

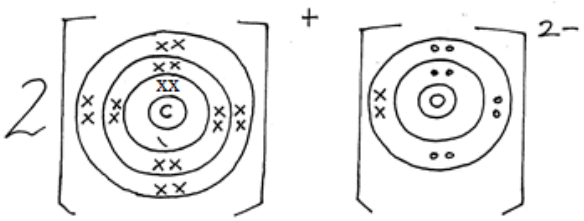
CHEMISTRY 233/2

PAPER 2

1. a) i) A- 2.8

B – 2.8.8

ii)



b) Shown on the grid between (B and D)

c) i) A is less reactive ✔1 than C, C has a larger atomic radius ✔1 hence loses its outermost electrons more easily.

ii) B has a smaller ✔1 atomic radius than A since B has stronger nucleus ✔1 charge.

ii) Oxide of G has a higher melting point than oxide of D ✔1 since G oxide is ionic and has strong ionic bonds ✔½ whereas oxide of D has a molecular structure with weak vanderwaal forces ✔½ between molecules.

d) Covalent bond ✔1

E and D share valence electrons ✔1 to form covalent bond.

e) E forms ion by gaining electrons ✔1. There exist repulsive forces ✔1 between the incoming electron and the existing electrons in E making the outer energy level bulge outward.

2. a) A – Brine/ concentrated sodium chloride. Reject sodium chloride

B – Carbon (iv) oxide

b) i) $NH_{3(g)} + H_2O_{(l)} + CO_{2(g)} \longrightarrow NH_4HCO_{3(aq)}$ ✔1

$NH_4HCO_{3(aq)} + NaCl_{(aq)} \longrightarrow NH_4Cl_{(aq)} + NaHCO_{3(s)}$ ✔1

iii) $Ca(OH)_{2(s)} + 2NH_4Cl_{(aq)} \longrightarrow CaCl_{2(aq)} + 2H_2O_{(l)} + 2NH_{3(aq)}$ ✔[1]

(award ½mk for correct equation without symbols)

c) i) Calcium chloride✔[1]

ii) I- Drying agent for gases/ as a drying agent✔[1] in the desiccators

II – In extraction of sodium from molten sodium chloride✔½. It lowers the melting point✔½ of NaCl from 801^0c to above 600^0c

d) i) Glass manufacturing industry

ii) Paper industry

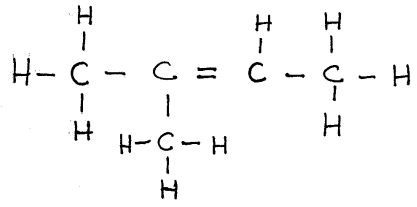
e) i) Efflorescence✔[1]

ii) –Decrease in mass✔½

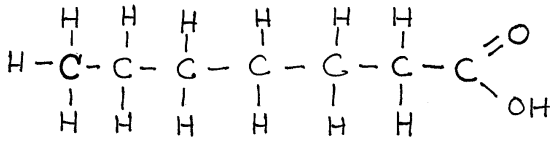
- Loss of crystalline nature✔½

3a)

i)



ii)



b) Heat the two substances separate and determine their boiling point✔[1]. Hexanol has a higher boiling point than methanol. ✔[1]

c) i) I – Substitution

II – Chloroethane

ii) <u>Condition</u>

Warming✔½

Concentrated sulphuric (vi) acid✔

<u>Reagent</u>

Propanoic acid ✔[1]

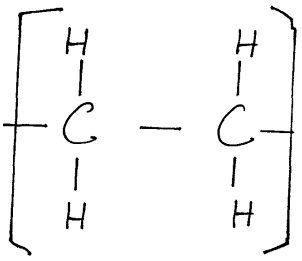
iii) CH_3CH_2ONa✔½ – Sodium Ethoxide✔½

iv) Hydrogen ✔[1]

Nickel catalyst✔½

Temperature 150^0c – 250^0c✔½

v) I



✔½

Polyethene✔½

II – Polythene bags✔1

$4.i) MnO_{2(s)} + 4HCl_{(l)} \rightarrow MnCl_2 + Cl_{2(g)} + 2H_2O_{(l)}$ ✔1

$ii) KMnO_{4(s)} / PbO_{2(s)}$ ✔1

iii) By passing it through concentrated sulphuric (vi) acid. ✔1

b) i) Aluminium chloride ✔1

$ii) 2Al_{(s)} + 3Cl_{2(g)} \rightarrow 2AlCl_{3(s)}$ ✔1

$iii) 2Al_{(s)} + 3Cl_{2(g)} \rightarrow 2AlCl_{3(s)}$ ✔1

$\frac{0.84g}{27} = 0.03111 moles$ ✔½

$Cl_{2(g)} volume used = 0.03111 x 3 = 0.09333 moles$ ✔½

Volume of chlorine = 0.09333 x 24000✔1

= 2240cm³

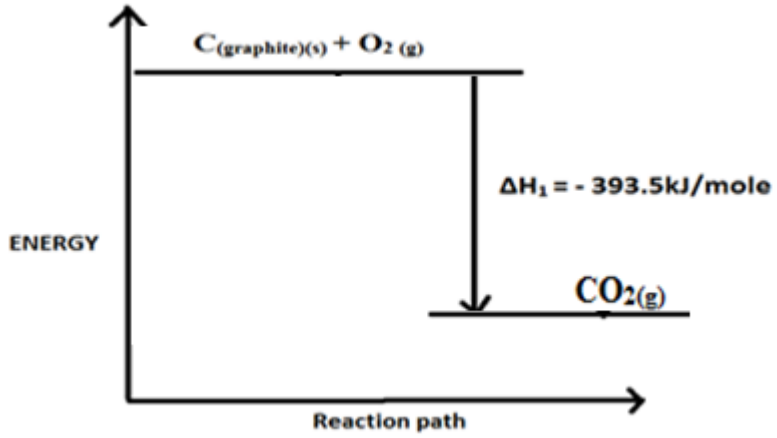
iv) – Calcium oxide prevents any moisture from outside since the $AlCl_3$ is deliquescent hence keeps combustion tube dry. ✔1

- Calcium oxide reacts with moisture forms calcium hydroxide that prevents chlorine from escaping to the atmosphere. ✔1

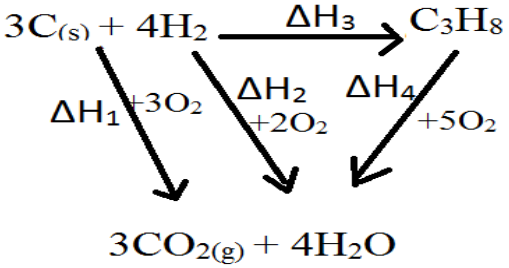
5.a) i) Energy or enthalpy change that occurs when a compound reacts completely with oxygen at standard conditions. ✔1

ii) I - Molar enthalpy of formation of propane. ✔1

II –



iii)



$\Delta H_1 + \Delta H_2 = \Delta H_3 + \Delta H_4$

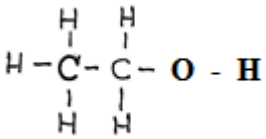
$3(-393.5) + 4(-285.8) = --103.7kJ/mole + \Delta H_4$ ✔½

$-1180.5 + -1143.2 = -103.7 + \Delta H_4$ ✔½

$-2323.7 + 103.7 = \Delta H_4$ ✔½

$\Delta H_4 = -2220kJ/mole$ ✔½

b)



$1(C-C) = -346kJ$
$5(C-H) = -2070kJ$
$1(C-O) = -360kJ$
$= -2776kJ$ ✔1

1 x -346kJ

5x -414 = - 2070kJ

1 x -360 = - 360Kj

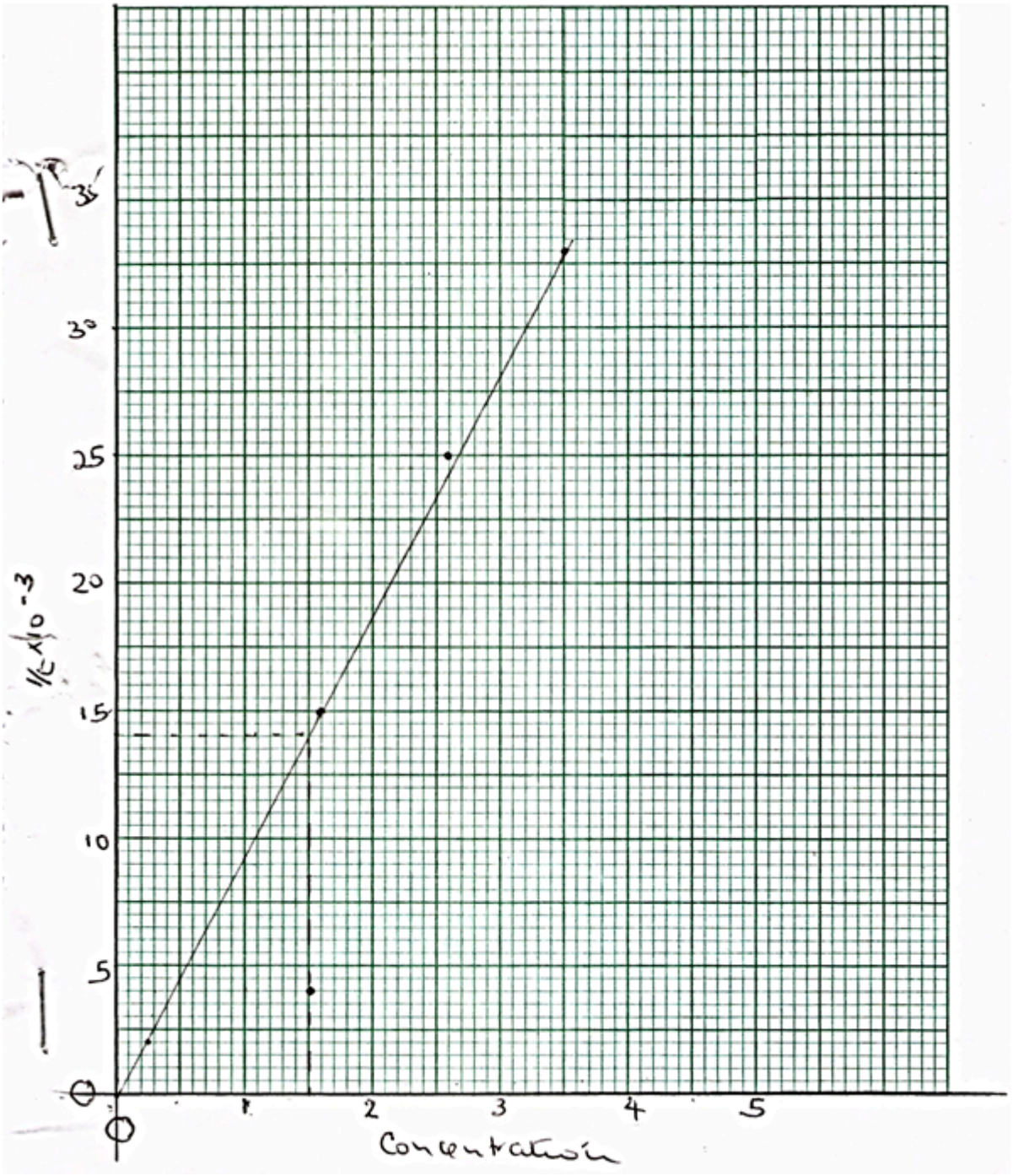
- 2776kJ + (O – H) = -3239 ✔1

O – H = -463kJ ✔1

6. a) i)

Acid concentration	0.25 M	1.5M	1.6M	2.6M	3.5M
Time in sec	500	250	67.5	40	30
$\frac{1}{time(s^{-1})}$	0.002	0.004	0.015	0.025	0.033

ii)



- Labelling – ✔½
- Scale – ✔½
- Plots – ✔1
- Line -✔1

NB: Straight line passing through the origin.

ii) 0.014✔ correct showing ✔½

Correct reading ✔½

iv) The rate of reaction increases with increase in concentration ✔1/ increased concentration increases the number of reacting particles and number of effective collisions✔1.

v) – Increased temperature (warm the mixture) ✔1

- Presence of a catalyst/ add crystals of $CuSO_4$

$vi) Zn_{(s)} + H_2SO_{4(aq)} \rightarrow ZnSO_{4(aq)} + H_{2(g)}$ ✔½

1: 1 1: 1✔½

65.4g of Zinc produce 22400cm³ of H_2 at STP✔½

0.26G of Zinc produces cm³ of H_2 at STP

$= \frac{0.26g \times 22400cm^3}{65.45}$ ✔1

= 89.05cm³✔½

7. a) Mixture of soluble and insoluble salt

b) i) Lead carbonate (Reject formula)

ii) Carbon (iv) oxide (Reject formula)

c) $PbCl_2$

d) Zn^{2+}, SO^{2-}_4 (Reject names)

e) $Ba^{2+}_{(aq)} + SO^{2-}_{4(aq)} \rightarrow BaSO_{4(s)}$ (Penalise ½for wrong state symbol)

f) $PbCO_3$ and $ZnSO_4$ (accept names tied to the two for 1mk)

g) i) $Zn(OH)_{2(s)} + 2OH^-_{(aq)} \rightarrow [Zn(OH)_4]^{2-}_{aq}$

(Reject fully if not balanced, penalise ½mk for wrong state symbols)

ii) Tetramine Zinc (ii) ions

h) Add excess lead (ii) oxide to dilute nitric (v) acid✔½.

Filter✔½ the unreacted lead (ii) oxide;

Add sodium sulphate✔½/ K_2SO_4/ H_2SO_4 to the filtrate

Filter✔½ to obtain lead (ii) sulphate as the residue

Wash the residue✔½ with distilled water

Dry the residue between filter paper✔½