

Name \_\_\_\_\_

H Chem, Period \_\_\_\_\_

***Ionic Nomenclature WS #1***

The following atoms or polyatomic ions will form ionic bonds. For each atom, give the ion it forms and then provide the ionic bond formula with an overall charge of zero.

| <b>Cation and Anion</b>    | <b>Positive Ion</b> | <b>Negative Ion</b> | <b>Formula</b>    |
|----------------------------|---------------------|---------------------|-------------------|
|                            |                     |                     |                   |
| potassium and iodine       | $K^{1+}$            | $I^{1-}$            | KI                |
| zinc and sulfur            | $Zn^{2+}$           | $S^{2-}$            | ZnS               |
| strontium and nitrite      | $Sr^{2+}$           | $NO_3^{1-}$         | $Sr(NO_2)_2$      |
| aluminum and acetate       | $Al^{3+}$           | $C_2H_3O_2^{1-}$    | $Al(C_2H_3O_2)_3$ |
| barium and nitrogen        | $Ba^{2+}$           | $N^{3-}$            | $Ba_3N_2$         |
| lithium and carbonate      | $Li^{1+}$           | $CO_3^{2-}$         | $Li_2CO_3$        |
| ammonium and bromine       | $NH_4^{1+}$         | $Br^-$              | $NH_4Br$          |
| silver and sulfate         | $Ag^{1+}$           | $SO_4^{2-}$         | $Ag_2SO_4$        |
| copper (II) and chlorine   | $Cu^{2+}$           | $Cl^{1-}$           | $CuCl_2$          |
| nickel (II) and phosphorus | $Ni^{2+}$           | $P^{3-}$            | $Ni_3P_2$         |
| lead (IV) and phosphite    | $Pb^{4+}$           | $PO_3^{3-}$         | $Pb_3(PO_3)_4$    |
| manganese (II) and oxygen  | $Mn^{2+}$           | $O^{2-}$            | MnO               |