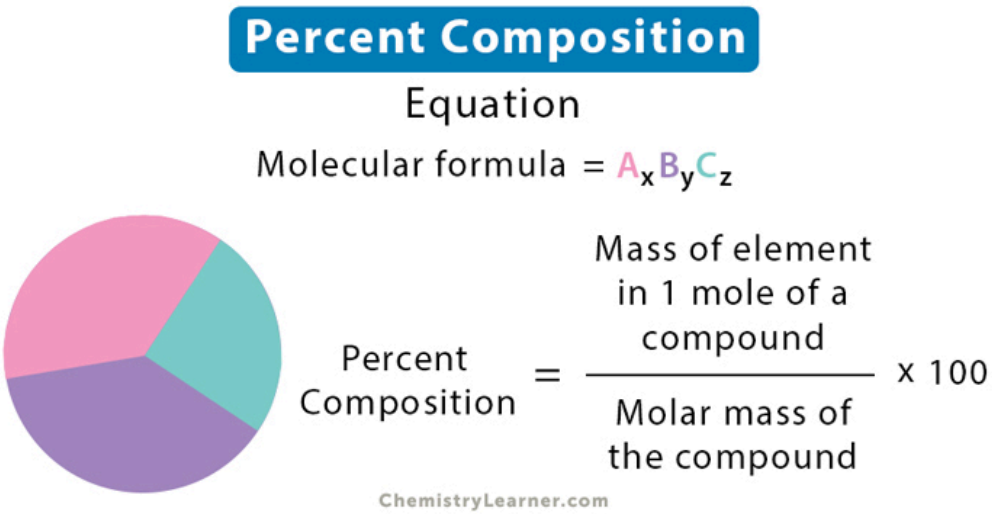
Honors Chemistry
BVHS
Unit: Stoichiometry

Chemical Composition-Percent Composition

Concepts

- Remember that a compound is a combination of two or more elements that are chemically bonded together.
- Molar Mass is the mass of one mole (6.022×10^{23} particles) of a substance.
- The percent composition tells you, by mass, what percentage of each element is present in the compound.
- It is simply doing a percentage calculation (part/whole then x 100)



Example: Determine the Percent Composition of Oxygen in water, H_2O.

Example A: Determine the Percent Composition of Carbon and Nitrogen in **$C_{17}H_{21}NO_4$**.

Extension: How many atoms of Carbon are in a 234g sample of $C_{17}H_{21}NO_4$?

Example B: Determine the Percent Composition of Oxygen and Hydrogen in **$C_{22}H_{28}N_2O$.**

Extension: How many atoms of Hydrogen are in a 234g sample of $C_{17}H_{21}NO_4$?