

-----KINETIC PARTICLE

THEORY-----

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Substance	Particle Model	Explanation
1A <i>Liquid mercury</i>	2C	As shown from the diagram, the particles are packed closely but there are still empty spaces between the particles. They are not arranged in an orderly manner, and they can slide over one another freely. Moreover, the attractive forces between the particles are strong. As such, 2C is the particle model of a liquid, thus it is matched with 1A.
1B <i>Common salt</i>	2E	Common salt is made of compounds and each compound's constituent elements are Sodium and Chloride. 2E illustrates this because in the diagram, the particles are very closely packed in an orderly manner, showing how the attractive forces between the particles are very strong. Also, the diagram has two types of particles, a smaller one (which is Na^+) and a bigger one (which is Cl^-). Na^+ is the smaller one because it only has two shells while Cl^- has 3 shells.
1C <i>Pure gold</i>	2A	Pure gold is a solid and as such the particles are packed very closely together, have an orderly arrangement and have a very strong attraction between them. As pure gold only has one type of element, Au, the diagram that fits is 2A.
1D <i>Evaporating water</i>	2F	The lower part of diagram 1D is the liquid state of water, because the particles shown are closely packed in a disorderly manner, and there are empty spaces between the particles. On the other hand, the upper part of the diagram shows water in its gaseous state, as the particles are very far from one another and they move about rapidly and freely in all directions.
1E <i>Bronze medal</i>	2B	Bronze is an alloy so it has two different elements. It is a solid so the particles are packed close together and have an orderly arrangement.
1F <i>Neon light</i>	2D	Neon is a gas and as such the particles do not have a strong attraction between them, making them far apart and randomly placed.

