

Strong Acids

Strong acids completely dissociate in water, forming H^+ and an anion. There are six strong acids. The others are considered to be weak acids. You should commit the [strong acids](#) to memory:

- HCl - hydrochloric acid
- HNO_3 - nitric acid
- H_2SO_4 - sulfuric acid
- HBr - [hydrobromic acid](#)
- HI - hydroiodic acid
- $HClO_4$ - [perchloric acid](#)

<http://chemistry.about.com/od/acidsbases/a/strengthacids.htm>

Weak Acids

A weak acid only partially dissociates in water to give H^+ and the anion. [Examples of weak acids](#) include [hydrofluoric acid](#), HF , and [acetic acid](#), CH_3COOH

Websites, images and videos to check out.....

[Acids explained](#)

[Concentrated vs Dilute; Strong vs Weak](#)

[The visual learner will love this one!!!!](#)

[Video - Look out, energetic Chemistry teacher!!!](#)