KAPSABET HIGH SCHOOL

1.(a) (i) Cracking√ 1

(ii) When the gas is burnt in air√ 1 it burns with a pale blue flame. √ 1

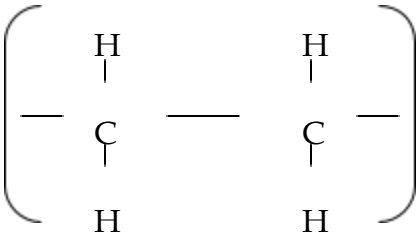
OR

Does not decolourize √ 1purple acidified <u>potassium manganate (VII).</u> √ 1

(iii) I. A. Ethane√ 1

II. B 1- Chloroethane√ 1

(iv)



n√ 1

(v) (i) Combustion√ 1

(ii) Dehydration√ 1

(vi) Conc. $H_2SO_{4\sqrt{1}}$

Temperature of 170^0C. √ 1

(b) (i) Pent-2-ene√ 1

(ii) Prop-1-yne. √ 1

1. a) (i) Mass of Mg is 20.36-19.52=0.84g√ ½

Mass of MgO is 20.92-19.52=1.40g

Mass of oxygen is 20.92-20.36=0.56g√ ½

% mass of Mg in MgO is $\frac{0.84}{1.40}\times 100\%$

=60%√ ½

%mass of O_2 in MgO is $\frac{0.56}{1.40}\times 100\%$

=40%√ ½

(ii)

Elements	Mg	O
%comp	60	40
R. A. M	24	16
Moles	60÷24 =2.5	40÷16 =2..5
Mole ratio	1√ 1	1√ 1

Empirical formula is thus MgO√ 1

b) (i) $2NaOH_{(aq)} + H_2SO_{4(aq)} \longrightarrow Na_2SO_{4(aq)} + 2H_2O_{(l)}$ √ 1

(ii)

I. Moles of H_2SO_4 in $20cm^3$ is $\frac{20\times0.25}{1000}$

=0.005moles√ 1

Mole ratio of base to acid is 2:1

Moles of NaOH is thus $\frac{2}{1}\times 0.005$

=0.01 moles of NaOH√ 1

II. Moles of NaOH in 1L is thus

$\frac{0.01\times1000}{50}$ √ 1

=0.2moles√ 1

III. Mass of NaOH in 1L is

0.2×40 =8g

Mass of NaCl in the mixture is 8.8-8= 0.8g√ 1

% mass of NaCl is $\frac{0.8}{8.8}\times100\%$

= 9.09%√ 1

2. a) (i) Bonds Broken are 4C-H= 4 x 413 = 1652

1C = C = 1 x 610 =610

1Br – Br = 1 x 193 = 193

Total energy absorbed = 2455 kJmol^{-1} √1

Bonds formed are 4C-H= 4 x 413 = 1652

2C = Br = 2 x 280 =560

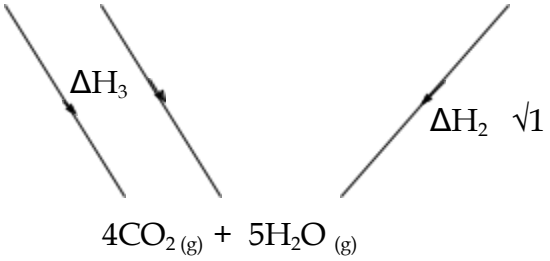
1C – C = 1 x 346 = 346

Total energy given out = 2558 kJmol^{-1} √1

ΔH = 2455 – 2558 = -103 **kJmol^{-1} √1**

(ii) Addition reaction √1

b) (i) $4C_{(s)} + 5H_{2(s)} \xrightarrow{\Delta H_1} C_4H_{10(g)}$



$\Delta H^{\theta}_{f}(C_4H_{10}) = \Delta H_3 - \Delta H_2$
$= 4(-393) + 5(-286) - (-2877)$ √1
$= -3002 + 2877$
$= -125\ kJmol^{-1}$ √1

(ii) $\Delta H + \Delta Hhyd$
$= 690 + -322 + -364$
$= 690 - 686$
$= +4\ kJ / mol$

3. a) √ 1
   i. Burrete √ 1
   ii. Pippete√ 1
   iii. Measuring cylinder

   b)
   i. Due to incomplete combustion, it produces white hot carbon particles that emittes a lot of light √ 1

   ii. It produces soot that makes apparatus dirty √ 1
   It does not produce much heat √ 1

   c)
   i. Nitrogen √ 1and oxygen√ 1
   ii. It can be separated by physical means √ 1
   Components of air are not chemically combined √ 1
   iii. Pass air through lime water ($Ca(OH)_2$√)√ 1 the lime water forms white precipitate indicating presence of carbon(IV)oxide √ 1

4. (a) Alkali metals √ 1
   (b) Electron arrangement 2.8.5 √1
   position: group V period 3√1

(c) The atom of R is larger √ ½ // has a larger atomic radius than the ion √ ½ This is because the ion of R is formed when the atom loses the electrons in the outermost energy level √ ½ therefore, the ion has one less energy level than the atom.√ ½

(d) (i) P_2W √ ½

(ii) TY_4 √ ½

(e) S has a higher √ ½ melting point than Q √ ½

This is because e S has more valence electrons in its metallic structure hence a stronger metallic bond √ ½ than Q √ ½

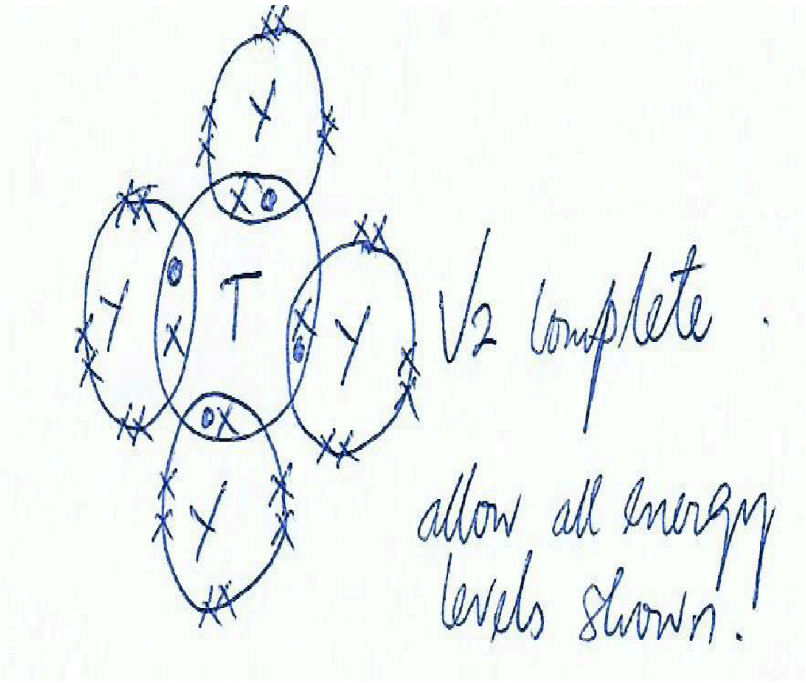
(f) M √1

It has a completely filled outermost energy level √ ½ and therefore, does not need to react with other elements to gain stability √ ½

(g) S has a higher electrical conductivity than Q√ 1.

S does not corrode easily like Q. √ 1

(h)



5. (a) A – Ammonia √1

B – Calcium oxide √1

(b) $CaO_{(s)} + H_2O_{(l)} \longrightarrow Ca(OH)_{2\,(aq)}$ √1

(c) Reaction is exothermic √1

(d) Filtration √1

(e) - Ammonia √1

- Carbon (IV) oxide √1

(f) (i) $C_{(s)} + 2H_2SO_{4(l)} \longrightarrow CO_{2(g)} + 2H_2O_{(l)} + 2SO_{2(g)}$ √1

(ii) Oxidising property √1

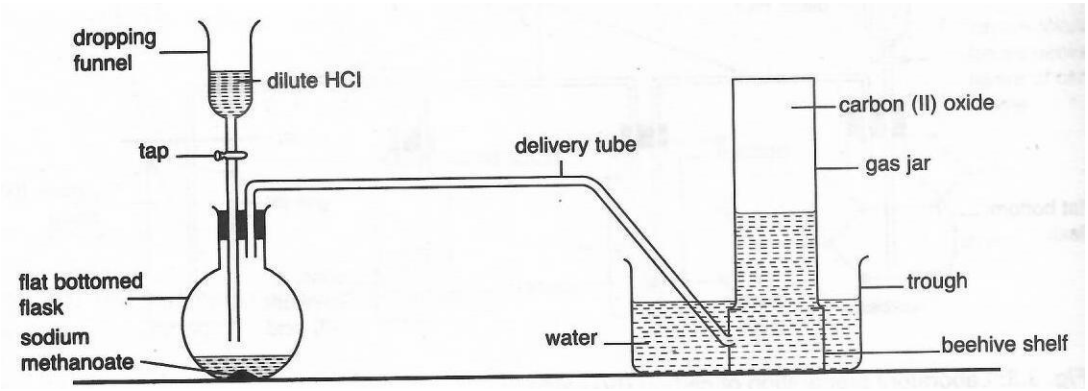
(g) - Manufacture of glass√½
- Softening of hard water √½
- Making of soaps and detergents
- For making sodium hydrogen carbonate used in baking soda and fire extinguishers

(Any 2 correct answers each ½ mk)

6. (a) (i)

Substance	Carbon (IV) oxide	Carbon (II) oxide
K	Dilute hydrochloric acid √½	Concentrated sulphuric (VI) acid √½
L	Marble chips or calcium carbonate √½	Sodium methanoate or ethanedioc acid (oxalic acid) √½

(ii)



Complete diagram = 1 mark, Labelling = 1 mark

(iii) Carbon (IV) oxide reacts with lime water / calcium hydroxide solution√½ to yield white precipitate while carbon (II) oxide does not. √½

Carbon (II) oxide burns √½with blue flame while carbon (IV) oxide does not burn. √½

(b) (i) $CO_{2(g)} + C_{(s)} \rightarrow 2CO_{(g)}$ √1

(ii) Reducing agent in extraction of some metals from their oxides. √1

(c) The bulb lights in set up I or conducts electricity while set up II does not. √ √1

In graphite, three out of four valence electrons of carbon atom are bonded leaving one delocalised√1 electron thus conducts electricity while in diamond, all the four valence electrons are bonded and is without delocalised electrons. √1