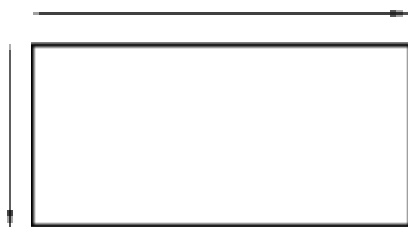


Periodic Trends for AP Chemistry

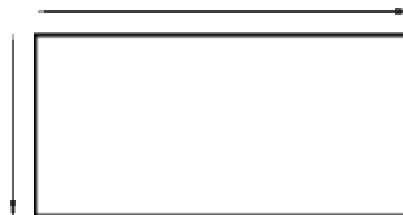
Explanations:

1. Added larger energy levels and shielding effect
2. Z_{eff} increases/nucleus becomes more positive (stronger)
3. Electron-electron repulsions

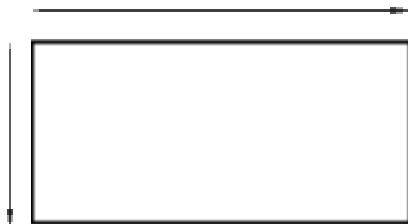
$Z_{\text{eff}} =$



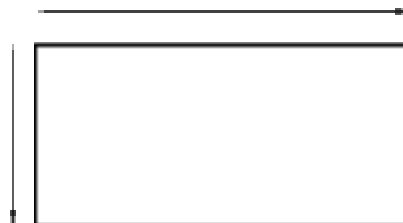
Ionization Energy =



Atomic Radius =



Electron Affinity =



Ionic Radius =



Electronegativity =

