

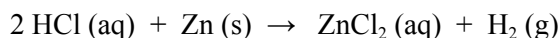
Name _____ Date: _____ Period: _____

Chemistry 09-13

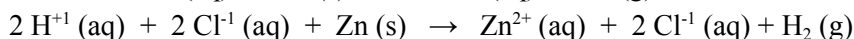
Ionic Equations

In many chemical reactions, some or all of the reactants and products may be ions, rather than molecules or atoms. Equations for single replacement and double replacement reactions are often written as ionic equations. Ionic equations can then be converted into net ionic equations by the omission of spectator ions.

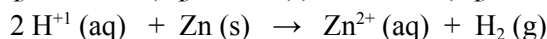
Single Replacement Reaction



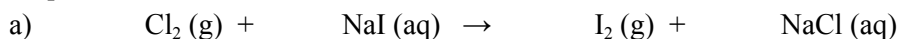
Ionic Equation



Net Ionic Equation

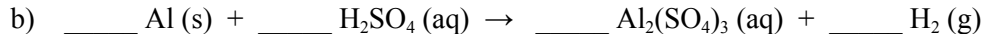


1) For each reaction below, balance the single replacement equation and convert it into ionic and net ionic equations.



b) Ionic equation _____

Net ionic equation _____



Ionic equation _____

Net ionic equation _____

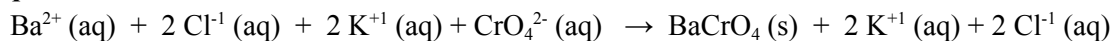


Ionic equation _____

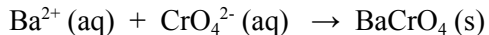
Net ionic equation _____



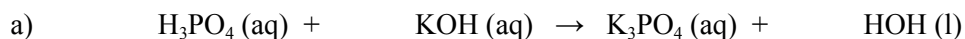
Ionic Equation



Net Ionic Equation

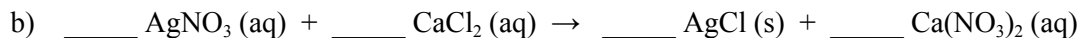


2) For each reaction below, balance the double replacement equation and convert it into ionic and net ionic equations.



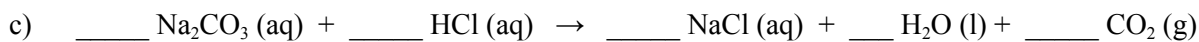
Ionic equation _____

Net ionic equation _____



Ionic equation _____

Net ionic equation _____



Ionic equation _____

Net ionic equation _____