Rb **O** Balance this equation: LiOH + H2C2O4 → Li2C2O4 + H2O. What are the coefficients?	2.1 g **A** Balance this equation: Li2S + H2O → H2S + LiOH What are the coefficients?	Li3N **S**	Li2S **D**
2 1 1 2 **V** Li2S + H2O → H2S + LiOH When 6.0 g Li2S is mixed with 4.5 g H2O, what is the limiting reactant?	1 2 1 2 **C** When LiOH and H2C2O4 react, what mass of H2O forms if 8.0 g H2C2O4 reacts?	1 2 3 1 **B**	Na **F**
H2O **E** An alkali metal, X, reacts with H2O to form hydrogen gas and the metal hydroxide. Write a balanced equation to show this. What are the coefficients?	3.2 g **T** When 1.0 g lithium nitride combines with 2.0 g water, forming nitrogen trihydride and lithium hydroxide what mass of excess reactant remains?	25% **H**	3 1 2 3 **K**
2 2 1 2 **R** When lithium nitride reacts with water, nitrogen trihydride and lithium hydroxide form. Write a balanced equation to show this.What are the coefficients?	0.4 g **I** An alkali metal, X, reacts with H2O to form hydrogen gas and the metal hydroxide. If 5.0 g of X forms 0.72 g H2, what is metal X?	56% **P**	Cs **W**
1 3 1 3 **R** Li2S + H2O → H2S + LiOH When 6.0 g Li2S is mixed with 4.5 g H2O and 4.0 g H2S forms, what is the percent yield?	Li **O** When LiOH and 8.0 g H2C2O4 reacts, and 2.5 g H2O forms, what is the percent yield?	14 g **G**	3 2 2 3 **L**
93% **E** When 1.0 g lithium nitride combines with 2.0 g water, forming nitrogen trihydride and lithium hydroxide what mass of lithium hydroxide should form?	78% **N** When 1.0 g lithium nitride combines with 2.0 g water, forming nitrogen trihydride and lithium hydroxide what is the limiting reactant?	4.5 g **U**	K **K**

The students knew they needed to review for their stoichiometry quest to avoid their teacher's _____.