

Name _____ Date _____ Period _____ AGS WB 15

Ionic Compounds and Formula Units

Directions Put the steps in order from first step to last step. The first step would be number 1. The last step would be number 4.

_____ Reduce the formula.

_____ Make the number in the positive charge the subscript for the anion. (Ignore charge)

_____ Write the cation and anion symbols with charges.

_____ Make the number in the negative charge into a subscript for the cation. (Ignore charge)

Directions Write the correct term on the line.

_____ 5. Ionic compounds are made of a cation and an anion, but are __.

_____ 6. Ionic formulas are reduced which is why they are called __.

Directions Write the formula unit for each ionic compound on the line.

_____ 7. sodium iodide

_____ 8. potassium sulfide

_____ 9. calcium iodide

_____ 10. copper (II) chloride

_____ 11. aluminum fluoride

_____ 12. iron (III) oxide

_____ 13. potassium chloride

_____ 14. potassium nitride

_____ 15. magnesium oxide