Honors Chemistry Name:
Worksheet 9.7 Date:

1. Using the "Activity Series" (HOT/NOT list) determine if these reactions will occur. If they occur, write ***the complete balanced equation***. IF NOT, write ***NO WAY JOSE!***

   a. silver nitrate + gold(II) □

   b. lithium chloride + fluorine □

   c. zinc acetate + hydrogen □

   d. chlorine + magnesium fluoride □

   e. potassium + zinc chromate □

2. Using the "Table of Solubilities" and your Periodic table, identify the phase (s, l, g, aq) of each substance in the following reactions. Complete the reactions as necessary.

   a) $2\ AlBr_{3(\ \ )} + 3\ Cl_{2(\ \ )} \rightarrow 2\ AlCl_{3(\ \ )} + 3\ Br_{2(\ \ )}$

   b) $Na_2CO_{3(\ \ )} + CaF_{2(\ \ )} \rightarrow CaCO_{3(\ \ )} + 2\ NaF_{(\ \ )}$

   c) $Ca_{(s)} + H_2O_{(\ \ )} \rightarrow$ _____$_{(\ \ )}$ + _____$_{(\ \ )}$

   d) ammonium sulfate $_{(\ \ )}$ + silver nitrate $_{(\ \ )}$ □

   _____$_{(\ \ )}$ + _____$_{(\ \ )}$

   e) potassium iodide $_{(\ \ )}$ + lead(II) nitrate $_{(\ \ )}$ □

   _____$_{(\ \ )}$ + _____$_{(\ \ )}$