

Name: _____ Date: _____ Pd: _____

Energy Levels

1st Read - Define the following vocabulary terms

Energy Level	
Atomic Model	
Valence Electrons	
Element	

2nd Read - Reading for Understanding

What Are Energy Levels

What did the author compare energy levels to in the text? Explain the similarities.	Electrons near the nucleus have higher energy than other electrons.
Explain how electrons can jump energy levels.	Explain how an electron gives off light energy.
Explain the relationship between energy levels to the amount of energy their electrons have.	

Energy Levels and Orbitals

Explain why electrons are always added to the lowest energy level.	
What determines the maximum number of electrons?	An atom has 18 electrons, how many orbitals does it have?

Energy Levels and Orbitals

Properties of an atom are determined by...	Explain how an atom is most stable
If an element is stable, explain why is it unlikely to react with other atoms?	In energy levels 2 - 7, how many electrons are needed to become stable?
Explain why Fluorine is reactive and how it would become stable?	
Explain why Lithium is reactive and how it would become stable?	