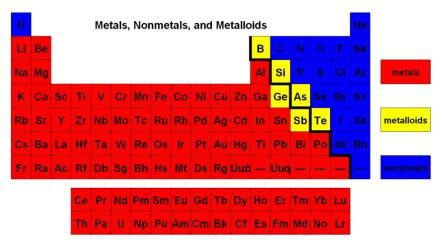
Ionic or Covalent

In each of the examples below you must identify the compound as being either IONIC or COVALENT. The type of bond that a compound forms is based on the types of elements bonded together. You must be able to identify elements as either METALS or NON-METALS. To do that you use their location on the periodic table:



Finally, to identify a compound as IONIC or COVALENT you simply need to use these rules:

METAL bonded to a NON-METAL \rightarrow IONIC NON-METAL bonded to a NON-METAL \rightarrow COVALENT

Metals do not bond to other metals!

Use the rules above to decide if each of the compound below is IONIC or COVALENT:

Name and Formula	Ionic or Covalent??
Sodium chloride, NaCl	
Calcium oxide, CaO	
Carbon monoxide, CO	
Nickel bromide, NiBr ₂	
Dinitrogen monoxide, N ₂ O	
Silver sulfate, Ag ₂ SO ₄	
Arsenic triiodide, Asl ₃	
Dibromine monoxide, Br ₂ O	
Molybdenum sulfide, Mo ₂ S ₇	
Sucrose, C ₁₁ H ₂₂ O ₁₁	
Tin chloride, SnCl ₂	
Sulfur hexafluoride, SF ₆	
Selenium phosphide, SeP	
Neodymium nitride, NdN	
Yttrium oxide, Y ₂ O ₃	
Nitrogen dioxide, NO ₂	
Rhodium selenide, RhSe ₂	
Sulfur dioxide, SO ₂	
Sodium sulfate, Na ₂ SO ₄	
Methane, CH₄	

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Name and Formula	Ionic or Covalent??
Sodium chloride, NaCl	IONIC
Calcium oxide, CaO	
Carbon monoxide, CO	COVALENT
Nickel bromide, NiBr ₂	
Dinitrogen monoxide, N₂O	COVALENT
Silver sulfate, Ag₂SO₄	
Arsenic triiodide, AsI ₃	COVALENT
Dibromine monoxide, Br ₂ O	
Molybdenum sulfide, Mo₂S ₇	IONIC
Sucrose, C ₁₁ H ₂₂ O ₁₁	
Tin chloride, SnCl ₂	IONIC
Sulfur hexafluoride, SF ₆	COVALENT
Selenium phosphide, SeP	
Neodymium nitride, NdN	IONIC
Yttrium oxide, Y ₂ O ₃	
Nitrogen dioxide, NO ₂	COVALENT
Rhodium selenide, RhSe ₂	
Sulfur dioxide, SO ₂	COVALENT
Sodium sulfate, Na ₂ SO ₄	
Methane, CH ₄	COVALENT