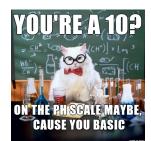
## **Introduction to Acid-Base Reactions**

# Graphic Organizer

# Acids **Ionization Reactions** $HCI(aq) + H_2O(I) \rightarrow$ $HCI(aq) \rightarrow$ Bases Metal Hydroxides $NaOH(aq) \rightarrow$ Nitrogen-Containing Molecules

 $NH_3(aq) + H_2O(I) \rightarrow$ 



### **Acid-Base Reactions**

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$$NaOH(aq) + HBr(aq) \rightarrow$$

$$H_2SO_4(aq) + KOH(aq) \rightarrow$$

## Carbonic Acid

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Predict the products of the following reactions

$$H_2SO_4(aq) + Al(OH)_3(aq) \rightarrow$$

$$NaHCO_3(aq) + HCI(aq) \rightarrow$$

$$Ba(OH)_2(aq) + HI(aq) \rightarrow$$

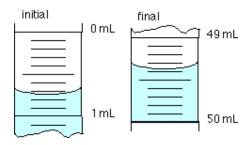
Titrations

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What is the concentration of 245 mL of HCl that is required to titrate 144 mL of 0.75M Ba(OH)<sub>2</sub>?