

Rules - Binary Ionic Compounds - [Helpful video](#)

1. The Cation (positive ion) is named first, the Anion second.

2. Monoatomic Cations take the element name

$\text{Na}^+ \rightarrow$ Sodium

$\text{Ca}^{2+} \rightarrow$ Calcium

3. Monoatomic Anions take the elements name and ends with "-ide"

$\text{Cl}^- \rightarrow$ Chloride

$\text{NaCl} \rightarrow$ Sodium Chloride

$\text{Li}_3\text{N} \rightarrow$ Lithium Nitride

Note** Greek prefixes are **not** used to indicate the number of atoms of each element in the formula unit for the compound

(e.g., Na_2O is named "sodium oxide" not "disodium oxide", or "disodium monoxide").

Formula	Name	Formula	Name
a. Li_2O	Lithium Oxide	n. K_2O	
b. CsCl		o. RbI	
c. MgS		p. Mg_3N_2	
d. CaBr_2		q. Ca_3P_2	
e. BaI_2		r. Al_2S_3	
f. KH		s. Ag_2O	
g. Na_2S		t. NaH	
h. AlN		u. CaO	
i. AgF		v. Na_2S	
j. ZnO		w. LiI	
k. CdI_2		x. CdS	
l. BaS		y. ZnCl_2	
m. LiCl		z. CsI	

**** Highlight to see your answer****