More stoichiometry practice

Use the balanced equation to answer the following questions:

$3\ C_2H_5O_2 + Al(OH)_3 \rightarrow Al(C_2H_4O_2)_3 + 3\ H_2O$
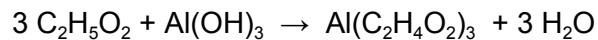
1. What is the mass of $Al(C_2H_4O_2)_3$ that can be formed from the reaction of 125 g of $C_2H_5O_2$ and 275 g of $Al(OH)_3$?
2. What is the limiting reagent? What is the excess reagent?
3. How many grams of H_2O are formed?
4. How much of the excess reagent is left over after the reaction is finished?
5. If 132 g of $Al(C_2H_4O_2)_3$ are formed during an experiment, what is the percent yield?