

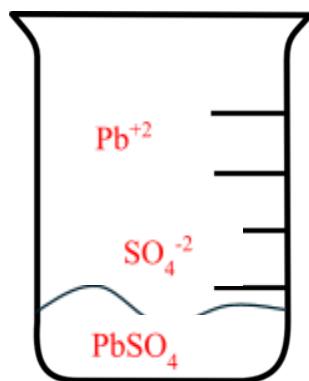
Name _____

H. Chem, Period _____

Electrolytic Solutions WS #2

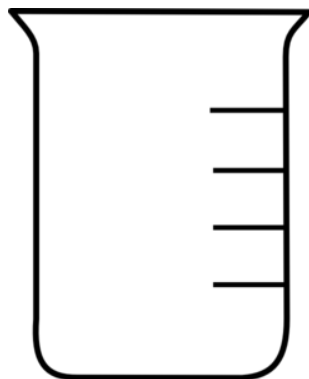
For each of the following, provide the name or the formula. Then classify each of the following solutes as a strong electrolyte (SE), weak electrolyte (WE), or nonelectrolyte (NE). Then, draw the resulting solution with water on a micro-molecular level.

Example: PbSO_4 lead (II) sulfate



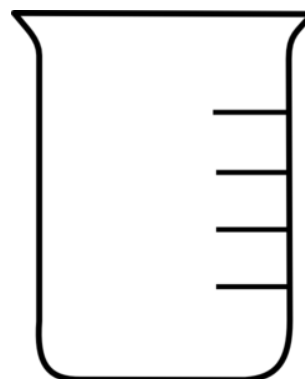
WE

1. SrI_2 _____

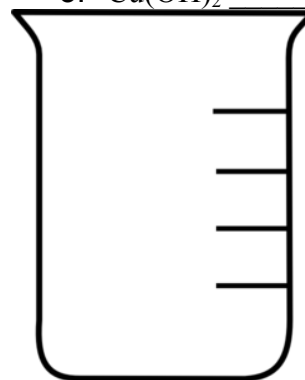


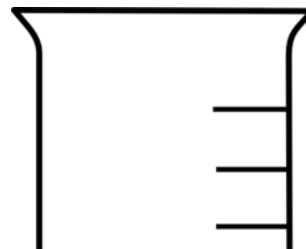
4. $\text{C}_6\text{H}_{12}\text{O}_6$ _____

2. $\text{Al}_2(\text{SO}_4)_3$ _____

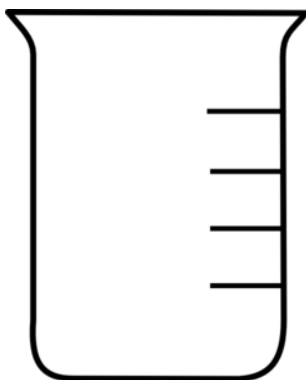


3. $\text{Cu}(\text{OH})_2$ _____

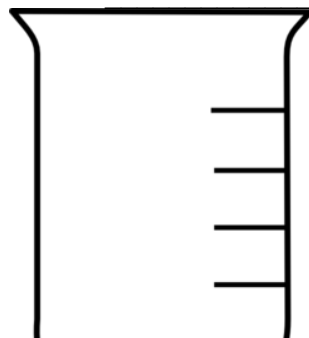




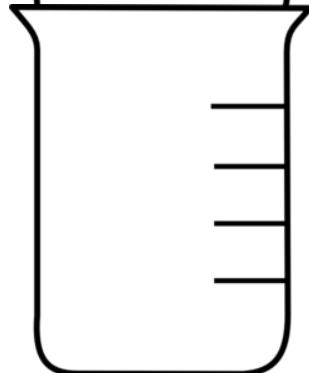
5. $\text{Ba}(\text{C}_2\text{H}_3\text{O}_2)_2$ _____



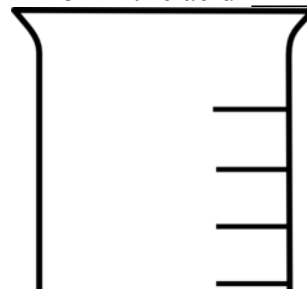
8. SrSO_4



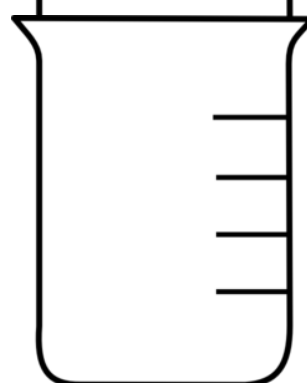
9. H_3PO_4 (aq)



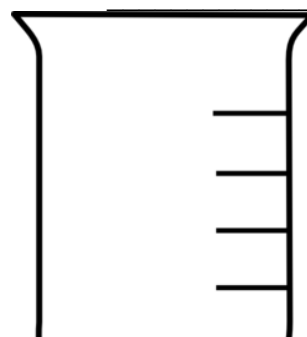
6. nitric acid



7. acetic acid



10. CO_2



11. HF (aq)

