Name_____ Honors Chemistry, Period____

Naming and Writing Covalent (Molecular) Compounds.

Part 1: Name the covalent or molecular compounds

a. CO carbon monoxide
b. CO_2 carbon dioxide
c. N_2O_3 dinitrogen trioxide
d. NP nitrogen monophosphide
e. SCl_2 sulfur dichloride
f. P_2O_5 diphosphorus pentoxide
g. NBr_3 nitrogen tribromide
h. CI_4 carbon tetraiodide
i. CCl_4 carbon tetrachloride
j. PF_5 phosphorus pentafluoride
k. SeF_2 selenium difluoride
l. $TeBr_2$ tellurium dibromide
m. P_2S_5 diphosphorus pentasulfide
n. C_3N_4 tricarbon tetranitride
o. CH_4 methane (carbon tetrahydride)
p. PH_3 phosphorus trihydride
q. SF_6 sulfur hexafluoride
r. IF_7 iodine heptafluoride

Part 2: Write the chemical formula

a. tetraphosphorus decoxide P_4O_{10}
b. tetraiodine nonoxide I_4O_9
c. carbon tetrafluoride CF_4
d. silicon dioxide SiO_2
e. dinitrogen trisulfide N_2S_3
f. phosphorus mononitride PN
g. carbon disulfide CS_2
h. dinitrogen trichloride N_2Cl_3

Name_____ Honors Chemistry, Period____

i. silicon tetrabromide $SiBr_4$
j. phosphorus triiodide PI_3
k. sulfur hexabromide SBr_6
l. oxygen difluoride OF_2
m. trisilicon tetranitride Si_3N_4
n. phosphorus pentachloride PCl_5
o. disulfur decafluoride S_2F_{10}
p. antimony pentabromide $SbBr_5$
q. dinitrogen pentoxide N_2O_5
r. sulfur trioxide SO_3