

## Molecular and Ionic Compound Structure and Properties

### 2.5 Lewis Diagrams

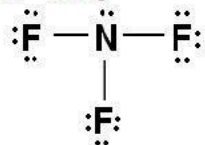
#### Worksheet Key

1) Draw Lewis diagrams for the following compounds:

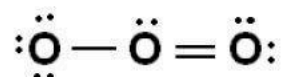
a.  $\text{CO}_2$



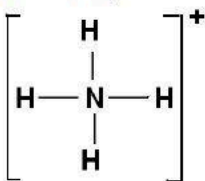
b.  $\text{NF}_3$



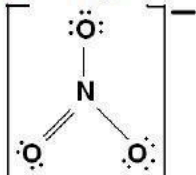
c.  $\text{O}_3$



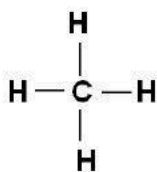
d.  $\text{NH}_4^+$



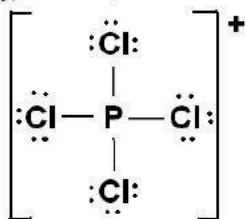
e.  $\text{NO}_3^-$



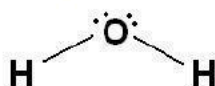
f.  $\text{CH}_4$



g.  $\text{PCl}_4^+$



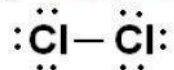
h.  $\text{H}_2\text{O}$



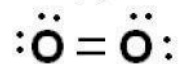
i. HCN



j.  $\text{Cl}_2$  (no central atom, make sure octets for both atoms are full)



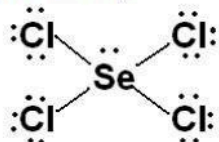
k.  $\text{O}_2$  (no central atom, make sure octets for both atoms are full)



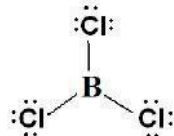
l.  $\text{N}_2$  (no central atom, make sure octets for both atoms are full)



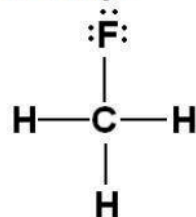
m.  $\text{SeCl}_4$



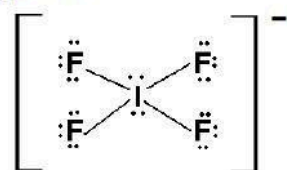
n.  $\text{BCl}_3$



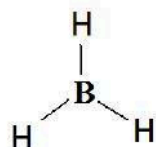
o.  $\text{CH}_3\text{F}$



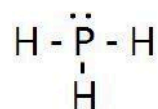
p.  $\text{IF}_4^-$



q.  $\text{BH}_3$



r.  $\text{PH}_3$



s. H<sub>2</sub>

