

Name: _____

H. Chem, Period: _____

Polyatomic Ionic Compound Nomenclature Worksheet

Part 1: Formula to Name

1. $\text{Zn}(\text{ClO}_2)_2$ _____
2. $\text{Mg}_3(\text{PO}_4)_2$ _____
3. ZnCr_2O_7 _____
4. CaCO_3 _____
5. KMnO_4 _____
6. H_2O_2 _____
7. $\text{AgC}_2\text{H}_3\text{O}_2$ _____
8. $\text{Cd}(\text{HSO}_4)_2$ _____
9. CuFO_3 _____
10. H_2SO_4 _____
11. $\text{Cr}_2(\text{SO}_3)_3$ _____
12. LiCN _____
13. Cu_3AsO_4 _____
14. NH_4OH _____
15. ZnSO_4 _____
16. Au_2CrO_4 _____
17. $\text{Sr}_3(\text{PO}_4)_2$ _____
18. AgMnO_4 _____
19. $\text{Co}(\text{HSO}_4)_2$ _____
20. $\text{HC}_2\text{H}_3\text{O}_2$ _____

Part 2: Name to Formula

1. sodium hydroxide _____
2. ammonium chromate _____
3. hydrogen iodate _____
4. lead (IV) oxalate _____
5. nickel (II) iodite _____
6. copper (I) sulfate _____
7. cobalt (III) hypochlorite _____
8. lithium acetate _____
9. mercury (I) dichromate _____
10. magnesium phosphate _____
11. mercury (II) hydroxide _____
12. zinc permanganate _____
13. manganese (VII) carbonate _____
14. potassium dichromate _____
15. iron (III) cyanide _____
16. lithium nitrite _____
17. beryllium thiosulfate _____
18. barium cyanide _____
19. ammonium permanganate _____
20. lead (IV) cyanide _____
21. sodium bicarbonate _____
22. beryllium acetate _____

Part 3: Mixed - Naming and Writing Formulas

1. barium sulfate _____
2. $\text{Mg}(\text{ClO}_4)_2$ _____
3. lead (II) acetate _____
4. LiClO_2 _____
5. nickel (II) hypochlorite _____
6. tin (IV) chlorate _____
7. Copper (I) hydroxide _____
8. manganese (IV) carbonate _____
9. copper (II) nitrite _____
10. $(\text{NH}_4)_2\text{C}_2\text{O}_4$ _____
11. CuCO_3 _____
12. iron (III) hydroxide _____
13. hydrogen cyanide _____
14. $\text{Cu}(\text{ClO}_3)_2$ _____
15. AgNO_3 _____
16. ammonium chloride _____
17. ammonium nitrate _____
18. $\text{Al}(\text{ClO})_3$ _____
19. NaCN _____
20. Na_3PO_4 _____