CLASSIC STUDY ONE OF THE BEST ONLINE STUDY PLAT FORM

N	ame of Student	Class	Subject	Board	T	Test #	Total N	Marks	Obtained Marks	
		10 th	Chemistry	(M.B)		Chapter:9 and	50			
						10				
Dat	te:		Objective	e		Teacher				
<u></u>						Remarks				
Q.1:										
1	In a chemical reaction substances that combines are called:									
a	Products	В			c	Reactants	d		Liquids	
2	The substances formed in a chemical reaction are called:									
a	Gases B Solids c Reactant d Products									
3	The reaction in which the products do not recombine to from reactants are called:									
a	Irreversible react	tion B	Exotherm	ic reaction	С	Spontaneous Reaction	d	Reversible reaction		
4	Reversible reaction	n are den	ted by:			reaction				
a	Single arrow	B		arrow	С	Three arrow	d		Four arrow	
5	An equilibrium is			unov		Timee untow	u		1 our unov	
a	Open system			system	С	Any system	d		All of these	
6	At equilibrium star			,		, ,	ı u		III OI HIODO	
a	Change	В	_		С	May or may not change	d		All of above	
7	At equilibrium star	tes a react	ion :			1 &-	<u> </u>			
a	Stops	В	1	ot stop	c	May or may not	d		None	
	1			1		stop				
8	With the passage of	of time rate	e of reverse re	action:		•				
a	Increased	В	Decre	eased	c	Constant	d	May i	ncrease or decreased	
9	With the passage of		_							
a	Increased	В			c	Decreased	d	May i	ncrease or decreased	
10	At equilibrium, the				reve					
a	Greater	В	Les	sser	c	Equal	d	NO	NE	
11	Active mass is exp									
a	[]	В)	С	{}	d		<>	
12	When $Q_c < K_c$ the r					T	1.1			
a	Forward	B	110,		С	At equilibrium	d		None	
13	The reaction has a					T .			NT.	
a	Small	В			CIZ	Large	d		None	
14	The reaction never				Ť		1 1		NI.	
1.5	Small The value of V do	B manda anl		ııum	С	Large	d		None	
15	The value of K _c de			aantratian	_	Doth a and L	L		None	
16	Temperature The word "acid " i				c	Both a and b	d		None	
16		s derived B		eet		C	d		Tagtalaga	
17	Bitter Which scientist pr				c o mo	Sour			Tasteless	
	Lavoisier	oved the p			e ma	nin constituent of all Humphrey	acias:		Newton	
18	Sometimes, the an				muol		u		INCWIOII	
a	Basidity	B			c	Coughness	d		Dizziness	
19	Acid turns litmus	<u>D</u>	ACI	uity		L Congilless	u		DIZZIIICSS	
a	Blue to red	b	Red to	n hlue	С	Blue to blue	d		Red to red	
20	The substance can				_	!	Į u		Acu to Icu	
a	Acidic	beliave as	Ba		c	Amphoteric	d		Neutral	
21	Conjugate base of		1 Da	510		1 mphotene	l u		rounai	
a	H ⁺	b	0	H-	c	H ₂ O	d		None	
22	The which of the f					<u> </u>			- · · · · · ·	
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_	Arrhenius	b	Bronsted- Lowry	С	Lewis	d	All of above				
23	Which of the following r					u	All of above				
23 a	AlCl ₃	h	CH ₄	Willi .	CCl ₄	d	CaO				
24		Which gas is produced when carbonates react with dilute acids									
a	H ₂	в	CO ₂	C C	NO ₂	d	N_2				
25	The reaction of acid with bases called										
a	Hydration	ь	Combustion	С	Nitration	d	Neutralization				
26	-	Which of the following acid is used for food preservation:									
a	Nitric acid	Ъ	Hydrochloride acid	С	Benzoic acid	d	Sulphuric acid				
27	Ants contain						2 to p to				
a	Butyric acid	b	Lactic acid	С	Formic acid	d	Nitric acid				
28	Which acid is used for et	ching		es:							
a	HCl	b	H_2SO_4	С	HNO ₃	d	None				
29	Which is mineral acid:		<u> </u>		,						
a	HCL	b	H_2SO_4	С	HNO ₃	d	All of above				
30	Apples contain which ac	cid									
a	Lactic acid	b	Malonic acid	С	Stearic acid	d	Uric acid				
31	The precipitate color of l	Pb (O	H) ₂ is:								
a	Blue	b	Green	c	White	d	Brown				
32	Which of the following h	nas bl	ue color precipitate:								
a	$Cu(OH)_2$	b	$Zn (OH)_2$	c	Fe(OH) ₃	d	$Ca(OH)_2$				
33	Which of the following i	s use	d in alkaline batteries								
a	КОН	b	NaOH	С	Al(OH) ₃	d	$Mg(OH)_2$				
34	At 25°C, the sum of pH a	_ ^	· •								
a	1	b	0	c	7	d	14				
35	It pOH is greater then 7 t				.	- 1					
a	Acidic	b	Basic	С	Neutral	d	None				
36	Which of the following i			,		1 . 1					
a	NaCl	b	KHSO ₄	<u> </u>	NH ₄ Cl	d	KCl				
37	Which of the following i					1 , 1	A 11				
a 20	Na ₂ SO ₄	b	NaClO ₃	С	Na_2SiO_3	d	All				
38	FeSO ₄ (NH ₄) ₂ SO ₄ .6H ₂ Ois			1	Г 1	1 1	0				
a 20	Potash alum	<u>b</u>	Mohr's salt	c	Ferric alum	d	Gypsum				
39	The formula of table salt		NaNO	Ι.	VNO	1	NaC1				
40	KCl The molar mass of H ₂ S	<u>b</u>	NaNO ₃	С	KNO ₃	d	NaCl				
	98g	U ₄ 18	98 amu	С	9.8 g	d	9.8 amu				
41	Which one of the follow:	ing el					9.6 amu				
a	Oxygen	b	Aluminum	c	Silicon	d d	Iron				
42	Gases are the lightest for						11011				
a	Mg cm ⁻³	b	G cm ⁻³	c	g dm ⁻³	d	Kg dm ⁻³				
43	Which one of the follow:	_			1 8	1	1 2-2				
a	Rubber	b	Plastic	С	Glass	d	Glucose				
44	Solid particles possess w	hich									
a	Rotational motions	b	Vibrational motion	С	Translational motion	d	Both c and b				
45	Transition elements are				1	1					
a	All gases	b	All metals	С	All non-metals	d	All metalloids				
46	Ÿ	Considering the electronic configuration of atoms which atom having given atomic number will be the most									
	stable one/		J 2 2- 2701112 WIII		62 , 411 4101						
a	6	b	8	С	10	d	12				
47	Which one of the follow			1		•					
a	BH_3	b	BF_3	С	N ₂	d	O_2				
		•	· · · · · · · · · · · · · · · · · · ·	•		•					

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48	The oxidation number of chromium in K ₂ Cr ₂ O ₇ is								
a	+2	b	+16	С	+7	d	+14		
49	Spontaneous chemical reaction take place in								
a	Electrolytic cell	b	Galvanic cell	С	Nelsons cell	d	Down's sell		
50	When U-325 breaks up it produces								
a	Electron	b	Neutrons	С	Protons	d	Energy		

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					Marks	
	10 th	Chemistry	(M.B)(F.B)	Chapter: 9 and 10	80	
Date :		Subjective		Teacher Remarks		

Section - B

Q.2:- Write short answers of the following question. (25 x 2 = 50)

- 1. Define a complete reaction?
- 2. What are the examples of natural for equilibrium phenomenon?
- 3. Define static equilibrium?
- 4. Define dynamic equilibrium?
- 5. Which basic properties does not change at equilibrium?
- 6. Why reversible reaction never complete?
- 7. Name two scientiest how put forward law of mass action and when?
- 8. State law of mass action?
- 9. What are the units of molar concentration?
- 10. Drive a K_c expression for a general reaction?
- 11. What do you mean by the extent of reaction?
- 12. How can sulphuric acid is prepare?
- 13. Write use of oxygen air?
- 14. Write the major components of atmosphere?
- 15. What is difference between Arrhenius base and Bronsted -Lawry base?
- 16. Prove that water is an amphoteric specie?
- 17. Define and give characteristics of a Lewis acid?
- 18. Why BF₃ behaves as a Lewis acid?
- 19. Define conjugate acid and conjugate base?
- 20. Give the uses of sulphuric acid?
- 21. Name the gas liberated when alkalies react with ammonium salts?
- 22. Write the names of some indicators?
- 23. What are the uses of pH?
- 24. What is differenece between "P" and "pH"?
- 25. How the basic salts turns into normal salts?

Section - C

Note: Write long answers of the following question. $(10 \times 3 = 30)$

- Q.3: (a) What is reversible reaction? Give one example.
 - (b) Discuss types of salts in detail.
- Q.4: (a) define chemical equilibrium state. Give possibilities at equilibrium state.
 - (b) Compare physical properties of acids and bases.
- Q.5: (a) Defines Acid and base with respect to Arrhenius concept. Also give their limitation.

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Write conjugate acid-base pairs of common species in a table from.

(b)

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