

CLASSIC STUDY ONE OF THE BEST ONLINE STUDY PLAT FORM

Name of Student	Class	Subject	Board	Test #	Total Marks	Obtained Marks
	10 th	Chemistry	(M.B)	Chapter :9 and 10	50	
Date : _____	Objective			Teacher Remarks		

Q.1:- Multiple choice Questions.

1	In a chemical reaction substances that combines are called:					
a	Products	B	Solids	c	Reactants	d Liquids
2	The substances formed in a chemical reaction are called:					
a	Gases	B	Solids	c	Reactant	d Products
3	The reaction in which the products do not recombine to from reactants are called:					
a	Irreversible reaction	B	Exothermic reaction	c	Spontaneous Reaction	d Reversible reaction
4	Reversible reaction are denoted by:					
a	Single arrow	B	Double arrow	c	Three arrow	d Four arrow
5	An equilibrium is achieved only in a:					
a	Open system	B	Close system	c	Any system	d All of these
6	At equilibrium state, the amount of reactants and products:					
a	Change	B	Do not change	c	May or may not change	d All of above
7	At equilibrium states a reaction :					
a	Stops	B	Dose not stop	c	May or may not stop	d None
8	With the passage of time rate of reverse reaction:					
a	Increased	B	Decreased	c	Constant	d May increase or decreased
9	With the passage of time rate of forward reaction:					
a	Increased	B	Constant	c	Decreased	d May increase or decreased
10	At equilibrium, the rate of forward reaction is -----to reverse reaction:					
a	Greater	B	Lesser	c	Equal	d NONE
11	Active mass is expressed as					
a	[]	B	()	c	{ }	d < >
12	When $Q_c < K_c$ the reaction goes in :					
a	Forward	B	Reverse	c	At equilibrium	d None
13	The reaction has almost gone complete when the value of K_c is					
a	Small	B	Medium	c	Large	d None
14	The reaction never goes to complete when the value of K_c is					
a	Small	B	Medium	c	Large	d None
15	The value of K_c depends only on					
a	Temperature	B	Initial concentration	c	Both a and b	d None
16	The word “acid “ is derived from a Latin word meaning .					
a	Bitter	B	Sweet	c	Sour	d Tasteless
17	Which scientist proved the presence of hydrogen as the main constituent of all acids:					
a	Lavoisier	B	Davy	c	Humphrey	d Newton
18	Sometimes, the amount of stomach acid becomes too much , which causes:					
a	Basidity	B	Acidity	c	Coughness	d Dizziness
19	Acid turns litmus					
a	Blue to red	b	Red to blue	c	Blue to blue	d Red to red
20	The substance can behave as an acid as well as a base is called					
a	Acidic	b	Basic	c	Amphoteric	d Neutral
21	Conjugate base of H_3O^+ is					
a	H^+	b	OH^-	c	H_2O	d None
22	The which of the following concept is only applicable for aqueous medium					

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a	Arrhenius	b	Bronsted- Lowry	c	Lewis	d	All of above
23	Which of the following molecule have central atom with incomplete octeti.						
a	AlCl ₃	b	CH ₄	c	CCl ₄	d	CaO
24	Which gas is produced when carbonates react with dilute acids						
a	H ₂	b	CO ₂	c	NO ₂	d	N ₂
25	The reaction of acid with bases called						
a	Hydration	b	Combustion	c	Nitration	d	Neutralization
26	Which of the following acid is used for food preservation:						
a	Nitric acid	b	Hydrochloride acid	c	Benzoic acid	d	Sulphuric acid
27	Ants contain						
a	Butyric acid	b	Lactic acid	c	Formic acid	d	Nitric acid
28	Which acid is used for etching designs on copper plates:						
a	HCl	b	H ₂ SO ₄	c	HNO ₃	d	None
29	Which is mineral acid:						
a	HCL	b	H ₂ SO ₄	c	HNO ₃	d	All of above
30	Apples contain which acid						
a	Lactic acid	b	Malonic acid	c	Stearic acid	d	Uric acid
31	The precipitate color of Pb (OH) ₂ is:						
a	Blue	b	Green	c	White	d	Brown
32	Which of the following has blue color precipitate:						
a	Cu(OH) ₂	b	Zn (OH) ₂	c	Fe(OH) ₃	d	Ca(OH) ₂
33	Which of the following is used in alkaline batteries						
a	KOH	b	NaOH	c	Al(OH) ₃	d	Mg(OH) ₂
34	At 25°C, the sum of pH and pOH is always equal to						
a	1	b	0	c	7	d	14
35	It pOH is greater then 7 than solution is						
a	Acidic	b	Basic	c	Neutral	d	None
36	Which of the following is an acidic salt						
a	NaCl	b	KHSO ₄	c	NH ₄ Cl	d	KCl
37	Which of the following is used in manufacture of explosives						
a	Na ₂ SO ₄	b	NaClO ₃	c	Na ₂ SiO ₃	d	All
38	FeSO ₄ (NH ₄) ₂ SO ₄ .6H ₂ Ois a formula of						
a	Potash alum	b	Mohr's salt	c	Ferric alum	d	Gypsum
39	The formula of table salt is						
a	KCl	b	NaNO ₃	c	KNO ₃	d	NaCl
40	The molar mass of H ₂ SO ₄ is						
a	98g	b	98 amu	c	9.8 g	d	9.8 amu
41	Which one of the following element is found in most abundance in the earth's crust?						
a	Oxygen	b	Aluminum	c	Silicon	d	Iron
42	Gases are the lightest form of matter and their densities are expressed in terms of						
a	Mg cm ⁻³	b	G cm ⁻³	c	g dm ⁻³	d	Kg dm ⁻³
43	Which one of the following is not amorphous ?						
a	Rubber	b	Plastic	c	Glass	d	Glucose
44	Solid particles possess which one of the following motion?						
a	Rotational motions	b	Vibrational motion	c	Translational motion	d	Both c and b
45	Transition elements are						
a	All gases	b	All metals	c	All non-metals	d	All metalloids
46	Considering the electronic configuration of atoms which atom having given atomic number will be the most stable one/						
a	6	b	8	c	10	d	12
47	Which one of the following is an electron deficient molecule?						
a	BH ₃	b	BF ₃	c	N ₂	d	O ₂

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48	The oxidation number of chromium in $K_2Cr_2O_7$ is						
a	+2	b	+16	c	+7	d	+14
49	Spontaneous chemical reaction take place in						
a	Electrolytic cell	b	Galvanic cell	c	Nelsons cell	d	Down's sell
50	When U-235 breaks up it produces						
a	Electron	b	Neutrons	c	Protons	d	Energy

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Section – B

Q.2:- Write short answers of the following question. (25 x 2 =50)

- Define a complete reaction?
- What are the examples of natural for equilibrium phenomenon?
- Define static equilibrium?
- Define dynamic equilibrium?
- Which basic properties does not change at equilibrium?
- Why reversible reaction never complete?
- Name two scientiest how put forward law of mass action and when?
- State law of mass action?
- What are the units of molar concentration?
- Drive a K_c expression for a general reaction?
- What do you mean by the extent of reaction ?
- How can sulphuric acid is prepare?
- Write use of oxygen air?
- Write the major components of atmosphere?
- What is difference between Arrhenius base and Bronsted –Lawry base?
- Prove that water is an amphoteric specie?
- Define and give characteristics of a Lewis acid?
- Why BF_3 behaves as a Lewis acid ?
- Define conjugate acid and conjugate base?
- Give the uses of sulphuric acid?
- Name the gas liberated when alkalies react with ammonium salts?
- Write the names of some indicators?
- What are the uses of pH?
- What is difference between “P” and “pH”?
- How the basic salts turns into normal salts?

Section - C

Note: Write long answers of the following question. (10 x 3 = 30)

- Q.3:** (a) What is reversible reaction? Give one example.
- (b) Discuss types of salts in detail.
- Q.4:** (a) define chemical equilibrium state. Give possibilities at equilibrium state.
- (b) Compare physical properties of acids and bases.
- Q.5:** (a) Defines Acid and base with respect to Arrhenius concept. Also give their limitation.

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(b) Write conjugate acid-base pairs of common species in a table from.