

## **QUESTION BANK**

## ORGANIC CHEMISTRY (3320501) /SEM-2

## **SUBJECTIVE QUESTIONS**

Sr. No. Question BTL CO  1 Write the types of structural isomerism 1 2 Define Homologous series. 1 3 Define isomerism. 1 4 Distinguish between organic and inorganic compound 1 5 Write homologous series of aldehyde 1 6 Write structural formula of following compounds (1)BHC (2)TNT (3)Propyne (4)oxalic acid. 1 7 Write short note on geometrical isomerism. 1 6 Write structural formula of following compounds. 1 8 (1) CH3CH-CH CH2CH3 (2) CH3OCH3 (3) CICH2COOH (4) CH3CH(OH)CH3 (5) CH3CCOOCH3 (6) CH3C OOC2H5 (6) CH3C OOC2H5 (7) CH3OCH3 (6) CH3C OOC2H5 (7) CH3CH3 (7) CH3OCH3 (7) CH3OCH	300	JECTIVE QUESTIONS	,	
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14 Define-Isomerism and explain position isomerism. 15 Give the general formula of alkane and alkene. 16 Give any two example of chain isomerism. 17 Explain position isomerism and functional group isomerism with example. 18 Give IUPAC name of the following compounds. (1) HCHO (2) CH3COOH (3) CH3COCH2CH3 (4) CH3COOCH2CH3  19 Write the structural formula of the following compounds. (1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid  20 Write structural formula of the following organic compounds. (a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane  21 Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  22 Give the types of isomerism 10 Give the structure of methanal and cyclo butane 11 Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane. 12 Write IUPAC name of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane. 12 Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3 10 Give the general formula of alkynes. 11 Define functional group. 12 Write structural formula of following compounds. 13 Cive the structural formula of following compounds. 14 Cive the general formula of following compounds. 15 Cive the general formula of following compounds. 16 Give the general formula of following compounds. 17 Define functional group. 18 Cive IUPAC nomenclature of following compounds. 19 Cive IUPAC nomenclature of following compounds. 10 CH2-CH2-CH2-CH2-CH2-CH2-CH2-CH2-CH2-CH2-				
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16Give any two example of chain isomerism.117Explain position isomerism and functional group isomerism with example.118Give IUPAC name of the following compounds. (1) HCHO (2) CH3COOH (3) CH3COCH2CH3 (4) CH3COOCH2CH3119Write the structural formula of the following compounds. (1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid120Write structural formula of the following organic compounds. (a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane121Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH122Give the types of isomerism123Give the types of isomerism124Write structure of methanal and cyclo butane124Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.125Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3126Give the general formula of alkynes.127Define functional group.128Write structural formula of following compounds. 1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol129Give IUPAC nomenclature of following compounds. 1. CH2=CH-CH=CH2,1	14			1
Explain position isomerism and functional group isomerism with example.  Give IUPAC name of the following compounds. (1) HCHO (2) CH3COOH (3) CH3COCH2CH3 (4) CH3COOCH2CH3  Write the structural formula of the following compounds. (1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid  Write structural formula of the following organic compounds. (a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane  Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  Give the types of isomerism  Give the structure of methanal and cyclo butane  Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1. CH2=CH-CH=CH2,	15			1
Give IUPAC name of the following compounds. (1) HCHO (2) CH3COOH (3) CH3COCH2CH3 (4) CH3COOCH2CH3  Write the structural formula of the following compounds. (1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid  Write structural formula of the following organic compounds. (a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane  Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  Give the types of isomerism  Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1 1. CH2=CH-CH=CH2,	16			1
19 (1) HCHO (2) CH3COOH (3) CH3COCH2CH3 (4) CH3COOCH2CH3  19 Write the structural formula of the following compounds. (1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid  20 Write structural formula of the following organic compounds. (a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane  21 Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  22 Give the types of isomerism  23 Give the structure of methanal and cyclo butane  24 Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  25 Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  26 Give the general formula of alkynes.  27 Define functional group.  28 Write structural formula of following compounds. 1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  29 Give IUPAC nomenclature of following compounds. 1 1 CH2=CH-CH=CH2,	17			1
(1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid  Write structural formula of the following organic compounds. (a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane  Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  Give the types of isomerism  Give the structure of methanal and cyclo butane  Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1 CH2=CH-CH=CH2,	18	• .		1
(1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid  Write structural formula of the following organic compounds. (a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane  Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  Give the types of isomerism  Give the structure of methanal and cyclo butane  Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1 1 CH2=CH-CH=CH2,	10	Write the structural formula of the following compounds.		1
(a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane   1	19	(1) Cyclopropane (2) Phenol (3) Methyl Benzene (4) Methanoic acid		1
(a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane  Write IUPAC name of the following organic compounds (a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  22 Give the types of isomerism 1 Give the structure of methanal and cyclo butane 1 Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1 CH2=CH-CH=CH2,	20	Write structural formula of the following organic compounds.		1
(a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  22 Give the types of isomerism  23 Give the structure of methanal and cyclo butane  24 Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  25 Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  26 Give the general formula of alkynes.  27 Define functional group.  28 Write structural formula of following compounds. 1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  29 Give IUPAC nomenclature of following compounds. 1 1 CH2=CH-CH=CH2,	20	(a) Methyl acetate (b) 2-propanone (c) 3-chloro pentanoic acid (d) Cyclopentane		1
(a) CH3COOH (b) CH3OCH3 (c) CH3CH2COOCH3 (d) CH3CH2OH  22 Give the types of isomerism  23 Give the structure of methanal and cyclo butane  24 Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  25 Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  26 Give the general formula of alkynes.  27 Define functional group.  28 Write structural formula of following compounds. 1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  29 Give IUPAC nomenclature of following compounds. 1 1 CH2=CH-CH=CH2,	21			1
Give the structure of methanal and cyclo butane  Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1. CH2=CH-CH=CH2,				
Write structural formula of the following organic compounds. (1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1 1-CH2=CH-CH=CH2,	22	Give the types of isomerism		1
(1) Ethyl acetate (2) 2,3- dimethyl hexane (3) 1,3-butadine (4) methoxy methane.  Write IUPAC name of the following organic compounds. (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  Give the general formula of alkynes.  Define functional group.  Write structural formula of following compounds. 1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1  Give IUPAC nomenclature of following compounds. 1	23	·		1
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25 (1) HCOOH (2) CH3-(CH2)6-CH3 (3) HCHO (4) CH3-CH2-COOCH3  26 Give the general formula of alkynes.  27 Define functional group.  28 Write structural formula of following compounds.  1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  29 Give IUPAC nomenclature of following compounds.  1 CH2=CH-CH=CH2,				
26Give the general formula of alkynes.127Define functional group.128Write structural formula of following compounds. 1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol129Give IUPAC nomenclature of following compounds.1. CH2=CH-CH=CH2,1	25			1
Define functional group.  Write structural formula of following compounds.  1 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1. CH2=CH-CH=CH2,	26			1
Write structural formula of following compounds.  1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol  Give IUPAC nomenclature of following compounds. 1. CH2=CH-CH=CH2,	27			1
1) 2-Butanone, 2) Ethanamide, 3) cyclo pentane, 4) 2-propene-1-ol Give IUPAC nomenclature of following compounds. 1. CH2=CH-CH=CH2,	2.0			_
Give IUPAC nomenclature of following compounds. 1. CH2=CH-CH=CH2,	28	- •		1
<sup>23</sup> 2. CH3CH2CH(CI)COOH 3. CH3CH(OH)CH3 4. CH3CH2CH2CHO 5. CH3COCH2COCH3	20			1
	29	2. CH3CH2CH(CI)COOH 3. CH3CH(OH)CH3 4. CH3CH2CH2CHO 5. CH3COCH2COCH3		1

1	Give the criteria for pure organic compound		2
2	Name the different methods for purification of liquids		2
3	Describe the method for purification of aniline.		2
4	Describe method used to determine boiling point of an organic compound		2
5	Write the boiling points of Aniline and nitrobenzene.		2
6	Define sublimation and give any two examples of sublimate compounds.		2
7	Explain crystallization.		2
8	Explain method for detection of melting point of organic solid.		2
9	Give the name of different methods for purification of organic liquids.		2
10	Describe the method for the purification of naphthalene		2
11	Write a short note of simple distillation method.		2
12	Give Name any two compounds that can be purified by sublimation method.		2
13	Explain Simple Distillation method to purify organic liquid.		2
14	Explain sublimation method for organic compound.		2
15	Which element are estimated by carius method?		2
16	Write boiling point of methanol and ethyl acetate.		2
17	Explain fractional distillation method for purification of organic liquid compounds with		7
17	figure.		2
18	Name the methods for purification of aniline and glycerin.		2
1	Tollen's reagent is prepared by adding into		3
2	Explain method for detection of carbon and hydrogen.		3
	Write the equations for the reactions involved in Lassaigne's test used for detection of		2
3	sulphur in given organic compound.		3
4	0.147 gm of an organic compound is heated with excess conc. HNO3 and AgNO3 gave		3
4	0.287 gm silver chloride. Find the percentage of chlorine in given compound.		3
5	Write about Lassaigne's test for detection of Nitrogen.		3
6	0.32 gm of an organic substance in a qualitative analysis gave 0.2334 gm of Barium		3
0	Sulphate. Calculate the percentage of sulphur.		э 
	0.20gm of an organic substance when heated with excess of strong Nitric acid and silver	]	
7	nitrate gave 0.3522gm of Silver iodide. Calculate the percentage of iodine in the		3
	compound.		
8	Explain Vacuum Distillation method.		3
9	Explain Kjeldahl's method for estimation of nitrogen.		3
10	Explain Duma's method for estimation of Nitrogen.		3
11	Explain Carius method for estimation of halogen.		3
12	Describe carius method for estimation of halides in organic compounds.		3
13	Explain lassaigne's test for the detection of nitrogen and Sulphur in an organic compound		3
	with equation.		
14	0.250 gm of an organic compound is heated with excess conc. HNO3 and AgNO3 gave		3
-7	0.4550 gm silver iodide. Find the percentage of iodine in given compound.		
15	Describe carius method for estimation of Sulphur in organic compound.		3
16	What is Lassaigne solution?		3
17	0.6 gm of an organic compounds give on combustion 0.150 gm of water and 0.350 gm of		3
1/	CO2, calculate percentage of carbon and hydrogen.		
18	0.45 g of an organic compounds give on combustion 0.300 g of water and 0.520 g of		3
10	carbon dioxide, calculate percentage of carbon and hydrogen.		<i>J</i>
19	0.301 gm of an organic compound is heated with excess conc. HNO3 and AgNO3 gave		3
10	0.282 gm silver bromide. Find the percentage of bromine in given compound.		,

2 Wri 3 (1) (3) 4 Wri 5 Wri 6 (1) (3) 7 Hov 8 Wri 10 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 2 Wri 21 Wri 22 Wri 23 Wri 2 Wri 3 Giv	· ,	4 4 4 4 4 4 4 4 4 4 4 4 4
2 Wri 3 (1) (3) 4 Wri 5 Wri 6 (1) (3) 7 Hov 8 Wri 10 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 2 Wri 21 Wri 22 Wri 23 Wri 2 Wri 3 Giv	rite the methods for preparation of ethyl alcohol  Ive the structure of products obtained during the following reactions  CH3CH2Cl + alc KOH = ?  C2H5OH + CH3COOH =?  (4)  CH2= CH2 + PdCl2+ H2O = ?  Tite the method of preparation, properties and uses of Acetone  Tite Williamsons synthesis of ether.  Ive the structure of products obtained during the following reactions  C2H5OH+PCl5 = ?  (2)CH3-CH=CH-CH3+O3[Zn] = ?  C6H5CH3 [FeCl3/Cl2] = ?  (4) 2CHI3 + 6Ag = ?  Ive will you distinguish between primary secondary and tertiary alcohol by oxidation.  Tite the preparation and uses of Ethanol.  Tite two uses of diethyl ether.  Tite physical properties and uses of Acetic Acid.  Tite the chemical properties and uses of Acetic acid.  Tite the preparation, physical properties and uses of methanol.  Tite the preparation, physical properties and uses of acetaldehyde.  Tite the physical and chemical properties of diethyl ether.  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid	4 4 4 4 4 4 4 4 4 4 4 4
Giv (3) 4 Wri 5 Wri 6 (1) (3) 7 Hov 8 Wri 10 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 L 21 Wri 22 Wri 23 Wri 3 Giv 3 Giv	ve the structure of products obtained during the following reactions  CH3CH2CI + alc KOH = ?  C2H5OH + CH3COOH =?  (4)  CH2= CH2 + PdCI2+ H2O = ?  Title the method of preparation, properties and uses of Acetone  Title Williamsons synthesis of ether.  Ve the structure of products obtained during the following reactions  C2H5OH+PCI5 =?  (2)CH3-CH=CH-CH3+O3[Zn] = ?  C6H5CH3 [FeCl3/Cl2] = ?  (4) 2CHI3 + 6Ag = ?  In will you distinguish between primary secondary and tertiary alcohol by oxidation.  Title the preparation and uses of Ethanol.  Title two uses of diethyl ether.  Title physical properties and uses of Acetic Acid.  Title the chemical properties and uses of Acetic acid.  Title the preparation, physical properties and uses of methanol.  Title the preparation, physical properties and uses of methanol.  Title the preparation, physical properties and uses of acetaldehyde.  Title the physical and chemical properties of diethyl ether.  Title the physical and chemical properties of oxalic acid  Title the physical and chemical properties of oxalic acid  Title the physical and chemical properties of oxalic acid  Title Wurtz reaction.	4 4 4 4 4 4 4 4 4 4
3 (1) (3) 4 Wri 5 Wri 6 (1) (3) 7 Hov 8 Wri 9 Wri 10 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 2. I 21 Wri 22 Wri 23 Wri 24 Giv 2 Wri 3 Giv	CH3CH2CI + alc KOH = ?  C2H5OH +CH3COOH =?  (4)  CH2= CH2 + PdCl2+ H2O = ?  Tite the method of preparation, properties and uses of Acetone  Tite Williamsons synthesis of ether.  We the structure of products obtained during the following reactions  C2H5OH+PCl5 = ?  (2)CH3-CH=CH-CH3+O3[Zn] = ?  C6H5CH3 [FeCl3/Cl2] = ?  (4) 2CHl3 + 6Ag = ?  We will you distinguish between primary secondary and tertiary alcohol by oxidation.  Tite the preparation and uses of Ethanol.  Tite two uses of diethyl ether.  Tite physical properties and uses of Acetic Acid.  Tite the chemical properties and uses of Acetic acid.  Tite the preparation, physical properties and uses of methanol.  Tite the preparation, physical properties and uses of acetaldehyde.  Tite the physical and chemical properties of diethyl ether.  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid  Tite Wurtz reaction.	4 4 4 4 4 4 4 4 4 4
(3) 4 Wri 5 Wri 6 (1) (3) 7 Hov 8 Wri 10 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 2. I 21 Wri 22 Wri 23 Wri 3 Giv	C2H5OH +CH3COOH =?  (4) CH2= CH2 + PdCl2+ H2O = ?  Title the method of preparation, properties and uses of Acetone  Title Williamsons synthesis of ether.  We the structure of products obtained during the following reactions  C2H5OH+PCl5 =?  (2)CH3-CH=CH-CH3+O3[Zn] = ?  C6H5CH3 [FeCl3/Cl2] =?  We will you distinguish between primary secondary and tertiary alcohol by oxidation.  Title the preparation and uses of Ethanol.  Title physical properties and uses of Acetic Acid.  Title physical properties and uses of Acetic Acid.  Title the chemical properties and uses of Acetic acid.  Title the preparation, physical properties and uses of methanol.  Title the preparation, physical properties and uses of acetaldehyde.  Title the physical and chemical properties of diethyl ether.  Title the physical and chemical properties of oxalic acid  Title the physical and chemical properties of oxalic acid	4 4 4 4 4 4 4 4 4 4
4 Wri 5 Wri 6 (1) (3) 7 Hov 8 Wri 9 Wri 10 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite the method of preparation, properties and uses of Acetone rite Williamsons synthesis of ether.  Ive the structure of products obtained during the following reactions  C2H5OH+PCI5 = ? (2)CH3-CH-CH3+O3[Zn] = ?  C6H5CH3 [FeCl3/Cl2] = ? (4) 2CHI3 + 6Ag = ?  Ive will you distinguish between primary secondary and tertiary alcohol by oxidation.  In the preparation and uses of Ethanol.  In the physical properties and uses of Acetic Acid.  In the physical properties and uses of ethyl acetate.  In the chemical properties and uses of Acetic acid.  In the preparation, physical properties and uses of methanol.  In the preparation, physical properties and uses of acetaldehyde.  In the physical and chemical properties of diethyl ether.  In the physical and chemical properties of oxalic acid  In the physical and chemical properties of oxalic acid  In the physical and chemical properties of oxalic acid  In the physical and chemical properties of oxalic acid  In the physical and chemical properties of oxalic acid  In the physical and chemical properties of oxalic acid  In the physical and chemical properties of oxalic acid	4 4 4 4 4 4 4 4 4
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6 (1) (3) 7 Hove (2) 8 Writh (3) 9 Writh (3) 10 Writh (3) 11 Writh (3) 12 Writh (3) 13 Writh (4) 15 Writh (4) 16 Writh (4) 17 Writh (4) 18 Diffe (4) 19 Writh (4) 20 2. If (4) 21 Writh (4) 22 Writh (4) 23 Writh (4) 24 Writh (4) 25 Writh (4) 26 Writh (4) 27 Writh (4) 28 Writh (4) 29 Writh (4) 20 Writh (4) 20 Writh (4) 21 Writh (4) 22 Writh (4) 23 Writh (4) 24 Writh (4) 25 Writh (4) 26 Writh (4) 27 Writh (4) 28 Writh (4) 29 Writh (4) 20 Writh (4) 20 Writh (4) 21 Writh (4) 22 Writh (4) 23 Writh (4) 24 Writh (4) 25 Writh (4) 26 Writh (4) 27 Writh (4) 28 Writh (4) 29 Writh (4) 20 Writh (4) 20 Writh (4) 20 Writh (4) 21 Writh (4) 22 Writh (4) 23 Writh (4) 24 Writh (4) 25 Writh (4) 26 Writh (4) 27 Writh (4) 28 Writh (4) 29 Writh (4) 20 Writh (4) 20 Writh (4) 20 Writh (4) 21 Writh (4) 22 Writh (4) 23 Writh (4) 24 Writh (4) 25 Writh (4) 26 Writh (4) 27 Writh (4) 27 Writh (4) 28 Writh (4) 29 Writh (4) 20 Writh (4) 21 Writh (4) 22 Writh (4) 23 Writh (4) 24 Writh (4) 25 Writh (4) 26 Writh (4) 27 Writh (4) 27 Writh (4) 28 Writh (4) 29 Writh (4) 20 Writh (4) 20 Writh (4) 20 Writh (4) 20 Writh (4) 21 Writh (4) 22 Writh (4) 23 Writh (4) 24 Writh (4) 25 Writh (4) 26 Writh (4) 27 Writh (4) 27 Writh (4) 28 Writh (4)	we the structure of products obtained during the following reactions  C2H5OH+PCI5 = ? (2)CH3-CH=CH-CH3+O3[Zn] = ?  C6H5CH3 [FeCl3/Cl2] = ? (4) 2CHI3 + 6Ag = ?  In will you distinguish between primary secondary and tertiary alcohol by oxidation.  In the preparation and uses of Ethanol.  In the physical properties and uses of Acetic Acid.  In the physical properties and uses of Acetic Acid.  In the chemical properties and uses of Acetic acid.  In the preparation, physical properties and uses of methanol.  In the preparation, physical properties and uses of acetaldehyde.  In the physical and chemical properties of diethyl ether.  In the physical and chemical properties of oxalic acid.  In the physical and chemical properties of oxalic acid.  In the physical and chemical properties of oxalic acid.  In the physical and chemical properties of oxalic acid.  In the physical and chemical properties of oxalic acid.  In the physical and chemical properties of oxalic acid.  In the physical and chemical properties of oxalic acid.	4 4 4 4 4 4 4 4 4
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(3) 7 Hove 8 Write 9 Write 10 Write 11 Write 12 Write 13 Write 15 Write 16 Write 17 Write 18 Diffe 19 Write 19 Write 12 Write 12 Write 11 Give 2 Write 13 Give 13 Give	C6H5CH3 [FeCl3/Cl2] = ?  W will you distinguish between primary secondary and tertiary alcohol by oxidation.  Tite the preparation and uses of Ethanol.  Tite two uses of diethyl ether.  Tite physical properties and uses of Acetic Acid.  Tite physical properties and uses of ethyl acetate.  Tite the chemical properties and uses of Acetic acid.  Tite the preparation, physical properties and uses of methanol.  Tite the preparation, physical properties and uses of acetaldehyde.  Tite the physical and chemical properties of diethyl ether.  Tite the physical and chemical properties of oxalic acid  Tite the physical and chemical properties of oxalic acid  Tite Wurtz reaction.	4 4 4 4 4 4 4 4
7 Hove 8 Write 9 Write 10 Write 11 Write 13 Write 15 Write 16 Write 17 Write 18 Diffe 19 Write 19 Write 12 Write 12 Write 11 Give 2 Write 3 Give	wwwill you distinguish between primary secondary and tertiary alcohol by oxidation.  The the preparation and uses of Ethanol.  The two uses of diethyl ether.  The physical properties and uses of Acetic Acid.  The physical properties and uses of ethyl acetate.  The the chemical properties and uses of Acetic acid.  The preparation, physical properties and uses of methanol.  The the preparation, physical properties and uses of acetaldehyde.  The physical and chemical properties of diethyl ether.  The the physical and chemical properties of oxalic acid  The Wurtz reaction.	4 4 4 4 4 4 4
8 Wri 9 Wri 10 Wri 11 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite the preparation and uses of Ethanol. rite two uses of diethyl ether. rite physical properties and uses of Acetic Acid. rite physical properties and uses of ethyl acetate. rite the chemical properties and uses of Acetic acid. rite the preparation, physical properties and uses of methanol. rite the preparation, physical properties and uses of acetaldehyde. rite the physical and chemical properties of diethyl ether. rite the physical and chemical properties of oxalic acid rite Wurtz reaction.	4 4 4 4 4 4 4
9 Wri 10 Wri 11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 3 Giv	rite two uses of diethyl ether.  rite physical properties and uses of Acetic Acid.  rite physical properties and uses of ethyl acetate.  rite the chemical properties and uses of Acetic acid.  rite the preparation, physical properties and uses of methanol.  rite the preparation, physical properties and uses of acetaldehyde.  rite the physical and chemical properties of diethyl ether.  rite the physical and chemical properties of oxalic acid  rite Wurtz reaction.	4 4 4 4 4 4
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11 Wri 12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite physical properties and uses of ethyl acetate. rite the chemical properties and uses of Acetic acid. rite the preparation, physical properties and uses of methanol. rite the preparation, physical properties and uses of acetaldehyde. rite the physical and chemical properties of diethyl ether. rite the physical and chemical properties of oxalic acid rite Wurtz reaction.	4 4 4
12 Wri 13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite the chemical properties and uses of Acetic acid. rite the preparation, physical properties and uses of methanol. rite the preparation, physical properties and uses of acetaldehyde. rite the physical and chemical properties of diethyl ether. rite the physical and chemical properties of oxalic acid rite Wurtz reaction.	4 4
13 Wri 14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite the preparation, physical properties and uses of methanol. rite the preparation, physical properties and uses of acetaldehyde. rite the physical and chemical properties of diethyl ether. rite the physical and chemical properties of oxalic acid rite Wurtz reaction.	4
14 Wri 15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite the preparation, physical properties and uses of acetaldehyde. rite the physical and chemical properties of diethyl ether. rite the physical and chemical properties of oxalic acid rite Wurtz reaction.	4
15 Wri 16 Wri 17 Wri 18 Diff 19 Wri 20 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite the physical and chemical properties of diethyl ether. rite the physical and chemical properties of oxalic acid rite Wurtz reaction.	
16 Wri 17 Wri 18 Diff 19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite the physical and chemical properties of oxalic acid rite Wurtz reaction.	4
17 Wri 18 Diff 19 Wri 20 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	rite Wurtz reaction.	
19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	ferentiate between aliphatic and aromatic compound.	4
19 Wri 20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv		4
20 Hov 2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	ite one method of preparation, property and use of ethylene	4
2. I 21 Wri 22 Wri 23 Wri 1 Giv 2 Wri 3 Giv	w will you effect following conversions 1. Acetaldehyde from acetylene	_
21 Wri 22 Wri 23 Wri  1 Giv 2 Wri 3 Giv	Ethyl acetate from acetic acid 3. Ethyl chloride from ethyl alcohol.	4
23 Wri	rite short note on Williamson's continuous etherification process	4
1 Giv 2 Wri 3 Giv	rite short note on Kolbe's electrolytic reaction.	4
2 Wri 3 Giv	rite test to distinguish between primary, secondary and tertiary alcohol.	4
2 Wri 3 Giv		
2 Wri 3 Giv	and the state of the land to the first of the land and the state of the land	
3 Giv	ve the structural formula of toluene and naphthalene	5
	rite the methods of preparations benzene	5
4 1 (-1)/	ve equation for following conversion (1) Toluene to benzaldehyde (2) Phenol to benzene.	5
	ve chlorination reaction of toluene.	5
	wwwill you convert benzene into toluene?	5
	rite the structures of Xylene and Naphthalene.	5
	rite the preparation and uses of Benzoic acid.	5
	rite the preparation and chemical properties of Phenol. Trite the structural formula of nitrobenzene and aniline.	5
	rite any two physical properties of oxalic acid.	5
	ve any two chemical properties of styrene.	5
	rite the method of preparation of Nitrobenzene.	5
	rite the method of preparation of Nitrobenzene.	5
		5
		5
	rite method of preparation and uses of Toluene	5
	rite physical and chemical properties of Benzene	
	rite physical and chemical properties of Benzene rite short note on Dow's process	[
	rite physical and chemical properties of Benzene rite short note on Dow's process rite the structure of Chloro benzene and Nitro benzene.	5
20 Exp	rite physical and chemical properties of Benzene rite short note on Dow's process	5 5 5

21	Write chemical properties and uses of benzoic acid		5
22	Write any two preparation method of salicylic acid.		5
23	Write the preparation, physical properties and uses of toluene.		5
24	Write the preparation, physical properties and uses of phenol.		5
25	Write physical and chemical properties of nitrobenzene.		5
26	Write physical and chemical properties of benzoic acid.		5
1	Write decarboxylation reaction		6
2	Give the name of reagents for nitration		6
3	Explain Friedel craft's alkylation reaction with suitable example		6
4	Define unit process. Explain. Name the reagent used in nitration.		6
5	Write preparation of nitrobenzene in the laboratory.		6
6	Give equation for following conversion (1) Benzene to aniline, (2) Ethanol to acetic acid.		6
7	Write a note on diazotization.		6
8	Write a short note on Friedel-Craft reaction.		6
0	Give equation for conversion of following. (1) Benzene to Benzaldehyde. (2) Nitrobenzene		
9	to aniline.		6
10	What are unit processes? Explain Sulphonation Unit process.		6
11	Write a short note on–(1) Tollen reagent Test (2) Reimer Tiemann reaction		6
12	What is Diazotization?		6
	Give chemical reaction for the following conversion (a) Benzene to Toluene (b) Toluene to		
13	Benzoic acid		6
14	Define the term Unit process. Explain nitration and halogenations unit process		6
15	Name reagents for nitration and sulphonation unit process.		6
1	What are carbohydrates? Give the classification of carbohydrate in detail.		7
2	What are carbohydrates give examples.		7
3	Give classification of Soaps and Detergent.		7
4	Write the classification of carbohydrates.		7
5	Explain classification of carbohydrate		7
6	Give the classification of soap and detergent.		7
7	Give the classification of soap and detergent with example.		7
,	dive the classification of soup and detergent with example.		<del>                                     </del>
1	What is Dye? Give example.		8
2	Define Chromophore and Auxochrome.		8
3	Explain difference between dye and color.		8
4	Define soap and detergent.		8
5	Give classification of dyes based on the type of application		8
6	Write classification of soaps and detergent with suitable example of each class.		8
7	Define Chromogen, Chromophore and Auxochrome		8
8	Give classification of dyes base on structure.		8
0	Give classification of tryes base on structure.		, °
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