Chemistry Day 6

Topic	Periodicity
NC Standards	Chm.1.3.2 Infer the physical properties (atomic radius, metallic and nonmetallic characteristics) of an element based on its position on the Periodic Table.
	Chm.1.3.3 Infer the atomic size, reactivity, electronegativity, and ionization energy of an element from its position on the Periodic Table.
Target Goals	 Define atomic radius, ionic radius, electronegativity, and ionization energy. Explain group and period trends for each. Compare the radius of a cation and anion to the radius of a neutral atom.
Learning Activity	 □ Watch the video on Periodicity (Bozeman) and answer the following questions. 1) What does the Aufbau principle determine? 2) Which law explains the attraction between electrons and the nucleus? 3) Define ionization energy 4) How does ionization energy change (increase or decrease) going across a period (left to right)? Why? 5) How does ionization energy change (increase or decrease) going down a group (top to bottom)? Why? 6) How does atomic radius change (increase or decrease) going across a period (left to right)? Why? 7) How does atomic radius change (increase or decrease) going down a group (top to bottom)? Why? 8) Define electronegativity. 9) Which group on the periodic table is the most stable? Why? 10) Describe one example of how knowing periodic trends is useful to chemists. □ Use the Periodicity Google Slides to review periodicity concepts and compare the radius of cations and anions There are 30 self-check quiz questions embedded in the slides. You will need a Chemistry Reference Sheet (Periodic Table) to answer the self-check questions as you move through the Google Slides.
Check for understanding.	Complete the following review activities to check for understanding of periodicity. You will need your Chemistry Reference Sheet (Periodic Table)
	Periodicity Quizizz
	Periodicity Pop-up Quia Quiz
	□ Periodicity Rags to Riches Review
	lons and Ionic Radius Quia Quiz