Lab 6 - Forming and Naming Ionic Compounds (Honors)

Introduction:
(Write 10 facts from Chapter 9.)

Materials:

2 well plates
0.2M KOH
0.2M Na_2SO_3
0.2M $CoCl_2$
0.2M $KMnO_4$
0.2M $NiCl_2$
0.2M Na_2CO_3
0.2M KI
0.2M $Pb(NO_3)_2$
0.2M $CuSO_4$
0.2M $Fe(NO_3)_3$
0.2M $AgNO_3$
0.2M $SrCl_2$
0.2M Na_3PO_4

Procedure:

1. Make sure your well plate is clean.
2. Add the chemicals to the well plate as follows in the data chart below.
3. Record the results in the data chart. You can tape the data chart into your lab notebook. Be sure to write in BLACK INK!

Data: (Fill in the chart with BLACK PEN, cut it out, and tape it into the lab notebook.)

	CoCl_2	**Pb**$(NO_3)_2$	**Cu**SO_4	**Fe**$(NO_3)_3$	**Ni**Cl_2	**Sr**Cl_2	**Ag**NO_3
Na_3**PO_4**							
K**OH**							
K**MnO_4**							
Na_2**CO_3**							

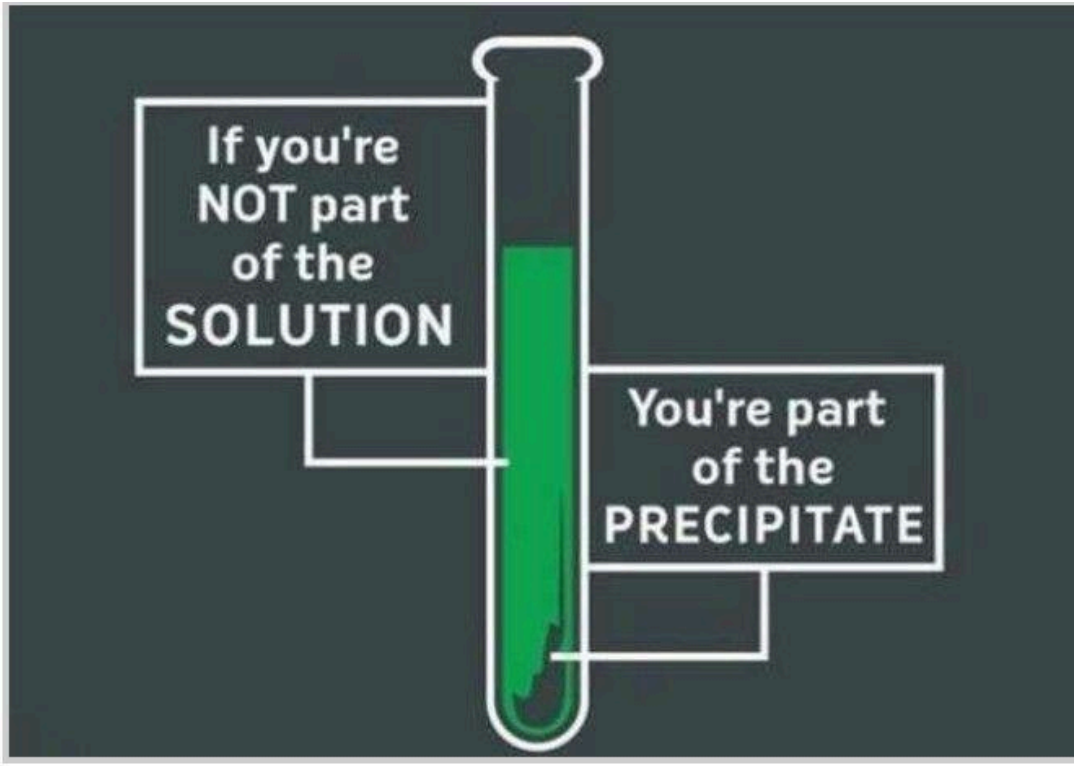
K**I**							
Na_2**SO_3**							

Questions: (Write the questions and answer them.)

1. What is the difference in the name of a compound that starts with a regular metal and one that starts with a transition metal?
2. What is a precipitate?
3. Name the following compound: Fe_2O_3.
4. Write the formula for the following compound: aluminum chlorate.
5. Predict whether the each of the following compounds is molecular or ionic:
   a. B_2H_6
   b. CH_3OH
   c. $LiNO_3$
   d. Sc_2O_3
   e. CsBr
   f. NOCl
   g. NF_3
   h. Ag_2SO_4

Conclusion:

(Write 3 sentences about any mistakes you made, anything that you learned, how the lab relates to real life, or other reactions that you have seen that are similar to this lab.)


If you're NOT part of the SOLUTION
You're part of the PRECIPITATE