Stoichiometry

Goes to the tune of "Thrift Shop" by Macklemore

Hey, Ms. Borders! Can you help us study?

What, what, what, what (4 times)

Study, Study, Better study! (4 times)

Chorus (Chapter 12)

Stoichiometry

The first step begins with converting to moles.

Then it's mole ratio, using coefficients.

Change to unit asked for.

Verse 1 (Chapter 10)

Chapter 10 begins with problems, they're called molar mass.

Add up the elements like we did in class.

The P T helps you achieve this goal.

Ms. Borders like, "Dang, the unit's grams per mole."

Number then, unit then, substance, follow what I do.

For atoms and molecules, use 6 point 0 2.

One mole equals molar mass, when you see g.

For liters one mole is twenty 2 point 4, see.

Next is percent composition!

The mass of the element, over the compound times one hundred, its fun.

The number of answers equals elements, then you're done.

And the unit is percent, man

or I take off half, cuz I can.

The empirical formula is the hardest to remember which

3 steps sometimes 4, 3 steps sometimes 4.

No, for real, percent to grams, step one is really easy.

Next change moles using molar mass.

Then divide by the smallest answer, and you pass.

Sometimes you get a point 5, sometimes you get a point 3.

Multiply all by 2, multiply all by 3.

The subscripts come from the answers you got.

Is molecular formula easy, or not?

M F over E F, gets a whole number.

The whole number times E F, "Aw, he got the answer!"

Chorus (times 2)

Verse 2 (Chapter 11)

What you know about crossing charges for formulas?

What you know about writing word equations?

You don't need to balance charges if you see prefixes.

Don't forget, the seven diatomics.

You start at element seven and make a big seven,

but you can't, forget hydrogen.

H, N, O, F, C l, I, and B r.

When you name, the roman numeral is the charge,

the charge, the charge, the charge, the CHARGE!!!

When you try to write word equations don't forget to write the, states of matter.

S is for solid, L is for liquid, states of matter.

G is gas, a q is aqueous, states of matter.

Synthesis is just like the marriage of two.

I'm like, "Decomposition is divorce, it's true."

Now single replacement has single on both sides, you know.

Combustion is the next one, CO_2 and H_2O.

Double displacement, switch positives.

In complete, separate all aqueous,

and balance charges.

When you predict the precipitates you need to use the solubility rules.

Solid, is insoluble, it's bionic.

Focus on precipitate for net ionic,

for net ionic.

Chorus

Bridge (Chapter 12)

Limiting reagent,

I will give you a hint.

Two 3 steps is the best,

limiting is the lowest.

Percent yield is the last,

the A is in the lab.

T is calculated.

A over T times one hundred.

Chorus