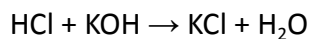


Titration Calculations

1. **10.0 cm³** of a solution of potassium hydroxide was titrated with a **0.10 M/dm³** solution of hydrochloric acid. **13.5 cm³** of the acid was required for neutralisation. Calculate the concentration of the potassium hydroxide solution.



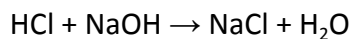
HCl

KOH

Concentration (mol/dm ³)	0.1	0.00135 / 0.01 =
Volume (cm ³)	13.5	10
Volume (dm ³)	13.5 / 1000 = 0.0135	10 / 1000 = 0.010
Moles	0.1 x 0.0135 = 0.00135	0.00135
Ratio	1	1

Answer:

2. 25.00cm³ of 0.100M/dm³ NaOH is needed to titrate 55.00cm³ of a solution of hydrochloric acid. Calculate the concentration of the acid.



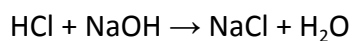
HCl

NaOH

Concentration (mol/dm ³)		
Volume (cm ³)		
Volume (dm ³)		
Moles		
Ratio		

Answer:

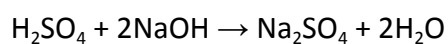
3. 6.25cm^3 of 0.125M/dm^3 NaOH is needed to titrate 25.00cm^3 of a solution of hydrochloric acid. Calculate the concentration of the acid.



Concentration (mol/dm^3)		
Volume (cm^3)		
Volume (dm^3)		
Moles		
Ratio		

Answer:

4. 25.00cm^3 of 0.2M/dm^3 NaOH is needed to titrate 12.50cm^3 of a solution of sulphuric acid. Calculate the concentration of the acid.



Concentration (mol/dm^3)		
Volume (cm^3)		
Volume (dm^3)		
Moles		
Ratio		

Answer: