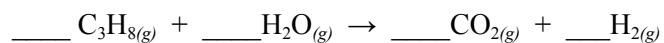
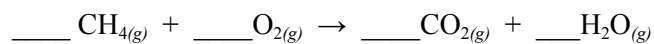
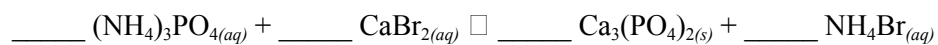
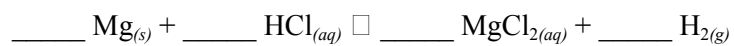


Chemistry 20 Unit Test Review

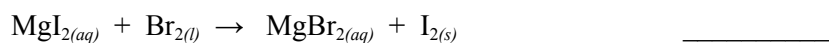
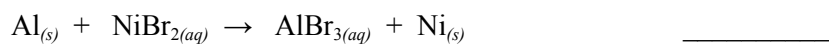
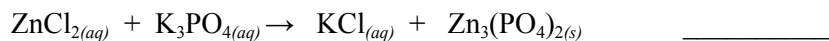
TOPIC 1: Naming

Name	Formula	Ionic, Molecular, Acid, Base
Lead (II) oxide		
Copper (I) sulfate		
	$\text{Pt}(\text{MnO}_4)_2$	
	HCl	
	NH_4Cl	
sodium carbonate		
	P_2O_5	
	K_3PO_4	
Calcium hydroxide		
	Fe_2O_3	
copper (I) chloride		
	H_2SO_3	
	CO_2	
chlorine		

TOPIC 2: Balancing



TOPIC 3: Reaction Types



TOPIC 4: Writing Reactions

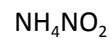
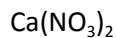
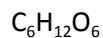
solid potassium chlorate decomposes into oxygen gas and solid potassium chloride

fluorine gas reacts with aqueous lead(IV) iodide

combustion of ethanol

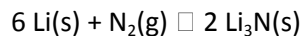
mercury(II) sulfide is decomposed into liquid mercury and solid sulfur

aqueous gold(III) bromide reacts with solid silver metal

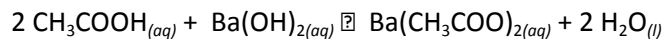
TOPIC 5: Molar Mass**TOPIC 6: Molar Mass Calculations**

What mass of calcium hydroxide is in 24 moles?

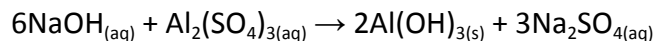
How many moles of calcium chloride are in 25.9g?

TOPIC 7: Stoichiometry

Determine the mass of N_2 needed to react with 0.536 moles of Li.



What mass of barium hydroxide is required to completely react with 10 mL of 0.88 mol/L acetic acid?



When 150.0 mL of sodium hydroxide, $\text{NaOH}_{(\text{aq})}$, reacts with aluminum sulfate, $\text{Al}_2(\text{SO}_4)_{3(\text{aq})}$, 0.0400 mol of aluminum hydroxide, $\text{Al}(\text{OH})_{3(\text{s})}$, precipitate. What is the concentration of the NaOH?

TOPIC 8: Titrations

10 mL of an unknown solution of NaOH is titrated 3 times with a 0.125M solution of HCl. What is the concentration of NaOH?

Titrant: HCl	Trial 1	Trial 2	Trial 3
initial volume (mL)	0.0	2.2	0.5
final volume (mL)	31.7	34.0	32.2
Volume Used (mL)			