Where is the new MS?????????? THERE IS A NEW ONE? SOMEONE PLEASE

There is another MS look at the second link on student room

Guyss please i need answers im panicking.

get to work my minions. I'm on phone so if someone sticks these in a table they're goated.

ETHANOL MOLES: 9.9x10-4 I think?
ETHANOL GDM3: I think it'd be somewhere forin the 110-115 range, say 113.
I got 111.9?

X,Y,Z elements: Group 1?
1st ionisation energy def: Energy required to ionise (remove a mole of electrons from) a mole of gaseous atoms

Reagent for that one cookie cutter mechanism: HCL I put LiAlH4? Dunno
I put AlCl3 and CH3Cl but think its wrong because it only had room for one / is this edexcel?

Question	Answer	marks
Define first ionization energy	Energy required to ionise (remove a mole of electrons from) a mole of gaseous atoms (to form 1 mol of cations)	2
X,Y,Z elements	Group 1 Because sharp increase from 1st to 2nd ionization energy So one valence electron	2
Which of XYZ has greatest atomic mass	X Because ionization energy decreases down the group X has lowest ionization energy so highest mass — IT was Z no?	3
What is the 2nd question		
(NH4)2CO3, K2CO3, K2SO4 testing	Idk for NH4+ NaOH, damp red litmus turns blue for NH4+	6

	Flame test for K+(can't do that experiment as wasn't provided with said have sample and water bath = which is for NaOH to test NH3) Acid and limewater for CO32- HCl+BaCl2 for SO42-	
Ethanol mols	9.9x10-4 BUT I GOT SOMETHING AORUND 1.08*10^-3 MOLES	5
Ethanol gdm^-3	110-115 (113/114?)	3
Question 8 i think?	N was 2???	
4 marker to show that only order of iodine was being investigated?	I calculated the moles for everything and then showed that the other two reactants were in excess meaning that moles and hence concentration of Iodine was the only thing affecting the order.	
NMR	I beg there was a septet somewhere?	
Recrystallisation one?	Soluble impurities reduced because the solvent causes the the desired product to form crystals and the soluble impurities remain in solution?	

What did we get for the ethanoic acid concentration?? OR WAS IT PROPANIOC ACID	I got 1. Something like it was over 1. I GOT 1.128	
Ph from the titration curve	ph=pka and ph was around 3.7 ? SAME AND THEN IN THE NEXT SECTION IT SAID TO CALCULATE THE KA VALUE WHICH I SIS 10^-3.6 OR 3.7	

Suggest an improvement to the student's technique (pouring acid from beaker to burette)	Lower the burette- same	
The student didnt remove the funnel, effect on n	Change in volume is lower than expected So volume of reacted naoh is greater So mol of acid is greater but mass is constant So Mr is lower So n is lower than expected Please how sure are we about n being lower? So many people said it was higher as well though	
Last question		