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Total No. of Printed Pages: [01]

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B.Pharmacy (Semester – 1st)
PHARMACEUTICAL ANALYSIS -I

Subject Code: BP102T

Paper ID: [17170102]

Time: 03 Hours

Maximum Marks: 75

Instruction for candidates:

1. Section A is compulsory. It consists of 10 parts of two marks each.
2. Section B consist of 9 questions of 5 marks each. The student has to attempt any 7 questions out of it.
3. Section C consist of 3 questions of 10 marks each. The student has to attempt any 2 questions out of it.

Section – A

(2 marks each)

Q1. Attempt the following:

- a) Define Primary Standard and secondary standard with examples
- b) Why pH control is important in complexometric titrations
- c) Define Co-Precipitation
- d) Define diffusion current
- e) How you will prepare 0.1 N sodium hydroxide in 500 mL of distilled water
- f) Define Aprotic solvents with examples
- g) Give the formulas for Muriatic acid and Oxalic acid
- h) Define Precision and accuracy
- i) Define significant figure with examples
- j) Define masking and demasking agents with suitable examples

Section – B

(5 marks each)

- Q2. Describe the limit test for sulphates as per 1985 and 1996
- Q3. Explain in detail Mohr's Method
- Q4. How you will prepare 0.1N Ceric ammonium sulphate. Describe its standardization procedure also
- Q5. Write a note on theories of acid base indicators
- Q6. What are metal ion indicators Discuss in brief.
- Q7. How you will prepare a) 0.1N Hydrochloric acid in 250 mL Distilled water and b) 0.1 N Sulphuric acid in 100 mL of distilled water.
- Q8. Give an account of Conductometric titrations
- Q9. Describe about Diazotization titrations and its significance.
- Q10. Write a short note on Iodometry

Section – C

(10 marks each)

- Q11. Give the principle of complexometric titrations and explain in detail
- Q12. Discuss in detail various sources and types of errors. Explain the methods how you will minimizing the errors.
- Q13. Define Polarography. Give its principle. Explain in detail construction and working of DME with neat labeled diagram.