

CONFIDENTIAL TO ALL SCHOOLS FOR CHEMISTRY TEACHERS

The information contained in this paper is to enable the Head of the school and teachers in charge of chemistry to make adequate preparation for this year's joint evaluation exams chemistry practical examinations.

NO ONE ELSE should have access to this paper or acquire knowledge of its content. Great care should be taken to ensure that the information contained herein **DOES NOT** reach candidates either directly or indirectly. The teacher in charge of chemistry **SHOULD NOT** perform any experiment in the same room as the candidates or make results of the experiment available to the candidates or give any other information related to the experiment to the candidates.

Requirements for candidates

In addition to the apparatus and fittings found in chemistry laboratory, each candidate will require the following

1. About 80cm³ of solution F
2. About 80cm³ of solution G
3. Burette
4. Pipette
5. Two conical flasks
6. Stop watch
7. Thermometer (- 10 to 110 °C)
8. 100 ml measuring cylinder
9. 100 ml beaker
10. Phenolphthalein
11. Distilled water
12. Solid H (3 cm length magnesium ribbon)
13. One label
14. Solid Q (1 g ammonium chloride)
15. 10 cm³ aqueous lead (II) nitrate
16. Red litmus paper
17. 6 test tubes
18. One boiling tube
19. Clean metallic spatula
20. Solution M (10 cm³ of lemon juice)
21. Source of heat
22. Acidified potassium manganate (VII)
23. Bromine water
24. Universal indicator and PH chart

25. 10 ml measuring cylinder.

N/B

Solution F is 0.5M sulphuric (VI) acid

Solution G is 0.5M sodium hydroxide