

Chemistry Day 8

| Topic | Metallic Bonds |
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| NC Standards | <p>Chm.1.2.1 Compare (qualitatively) the relative strengths of ionic, covalent, and metallic bonds.</p> <p>Chm.1.2.2 Infer the type of bond and chemical formula formed between atoms.</p> |
| Target Goals | <ul style="list-style-type: none"> <input type="checkbox"/> Explain how a metallic bond forms. <input type="checkbox"/> Explain how the properties of metals are related to metallic bonds. |
| Learning Activity | <ul style="list-style-type: none"> <input type="checkbox"/> Watch the Socratica Metallic Bonds Video and answer the questions below. <ol style="list-style-type: none"> Name 5 characteristics of metals. Are metals crystalline OR amorphous? Explain. In metals, the valence electrons are free to move among all metal cations. What term describes this phenomenon? True or False? Metallic bonds are much stronger than ionic and covalent bonds. Explain why metals are shiny. What determines the color of metals? Why are metals good conductors of heat? Explain! Why are metals good conductors of electricity? Explain! Are ionic compounds malleable? Explain why. Define malleable. Are metals malleable? Explain why. What is the most malleable metal? Define ductile. What is the most ductile metal? What is the only metal which is a liquid at room temperature? <input type="checkbox"/> Review metallic bonding by looking at the Metallic Bonds Google Slides |
| Check for understanding. | <p>Complete the following review activities to check for understanding of metallic bonds. You may need your Chemistry Reference Sheet (Periodic Table)</p> <ul style="list-style-type: none"> <input type="checkbox"/> Metallic Bonds Quizizz |