Chapter 4 Practice Quiz B
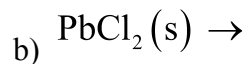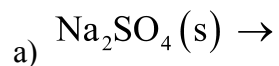
Name _____

2pts 1.   Complete the dissociation reactions for the following electrolytes in water:

a) $Na_2SO_4(s) \rightarrow$

b) $PbCl_2(s) \rightarrow$

2pts 2.   Complete the following with the correct symbol or name (correct spelling required).

a) acetic acid   _____

b) iron(III)periodate   _____

c) $NaH(SO_3)$ _____

d) $HClO_2$ _____

2pts 3. What is the molarity of a solution prepared by adding 5.00 mL of a 6.00 M NaCl solution into a 250 mL volumetric flask and diluting to the mark with water? (Ans: 0.12 M)

1pt 4.   What happens at the endpoint of a titration?

3pts 5. How many grams of precipitate could be formed when 1.25 L of 0.625 M $SnCl_2$ react with excess $(NH_4)_2CO_3$?   (Ans: 1.40 x 10^2 g)

$$SnCl_2(aq) + (NH_4)_2 CO_3(aq) \rightarrow 2NH_4Cl(aq) + SnCO_3(s)$$