Rules for Assigning Oxidation Number

NEET UG Chemistry — Quick Revision Notes

The oxidation number (O.N.) is the apparent charge assigned to an atom in a compound or ion using a standard set of rules.

1. Important rules

1

Free elements: O.N. = 0

Examples: $Na,\ Fe,\ H_2,\ O_2,\ Cl_2,\ S_8$.

All atoms in an element in its free state have oxidation number zero.

2

Monatomic ions: O.N. = ionic charge

Examples: $Na^{+} = +1,\ Mg^{2+} = +2,\ Cl^{-} = -1$.