

# Lewis Dot Structures Practice

## Multiple Bonds ONLY

Draw a Lewis Dot Structure for each molecule below. **All** of these structures require **MULTIPLE** (double and triple) bonds.

Some of these compounds can form multiple isomers. **Identify** which compounds can have multiple correct Lewis Dot Structures, and **draw 2 different isomers** for each of them!

I recommend doing this work on a separate document so you have the needed space. **If you have a stylus** do the work electronically. **If you do NOT**, do the work on a sheet of paper.

1.  $\text{O}_2$
2.  $\text{N}_2$
3.  $\text{C}_2\text{H}_4$
4.  $\text{C}_2\text{H}_2$
5.  $\text{C}_3\text{H}_6\text{O}$
6.  $\text{N}_2\text{H}_2$
7.  $\text{C}_4\text{H}_6$
8.  $\text{C}_4\text{H}_4\text{Cl}_2$
9.  $\text{C}_6\text{H}_6$
10.  $\text{C}_5\text{H}_{10}\text{O}_3$

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