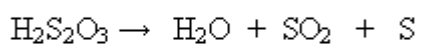
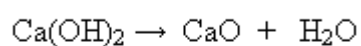
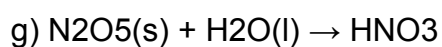
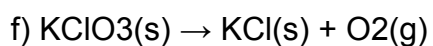
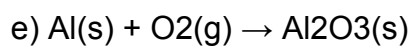
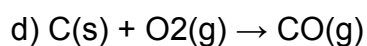
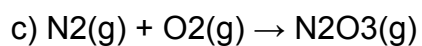
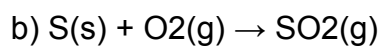
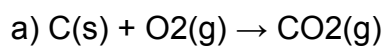
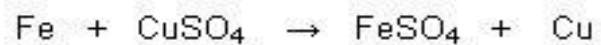
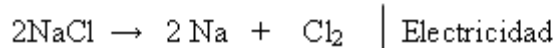
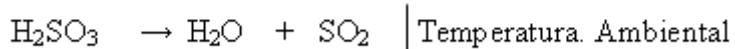


EXAMPLES OF CHEMICAL REACTIONS

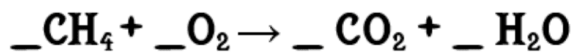


Calor*



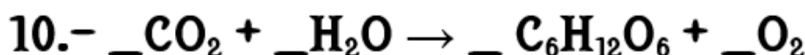
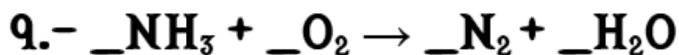
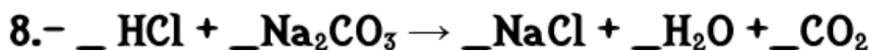
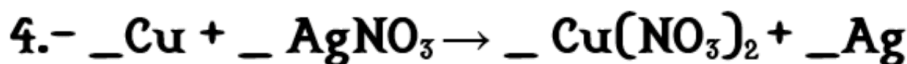
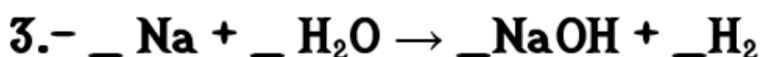
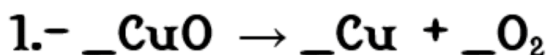
Balanceo de Ecuaciones Químicas

1.- Realiza el balance de la ecuación química siguiente por el método de tanteo. En función de la ecuación balanceada completa el cuadro inferior.



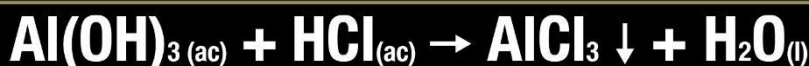
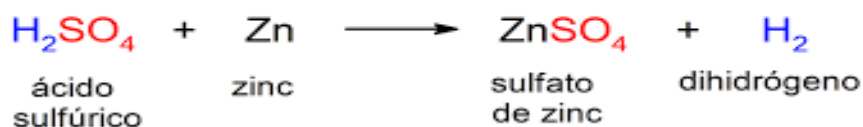
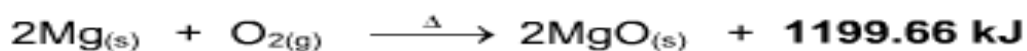
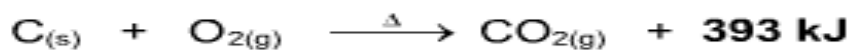
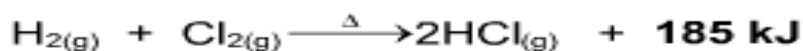
Número total de átomos de hidrógeno en los reactantes	
Elemento con el menor número de átomos.	
Número total de átomos de oxígeno en los productos	

2.- Balancear por tanteo las ecuaciones químicas siguientes y has tus anotaciones:



Carbono **Hidrogeno** **Oxigeno**





[]	$\text{Fe} + \text{S} \rightarrow \text{FeS}$	1	Reacción de descomposición
[]	$\text{Cl}_{2(g)} + \text{HI}_{(ac)} \rightarrow \text{HCl}_{(ac)} + \text{I}_2 \uparrow$		Reacción de combustión
[]	$\text{CaO}_{(s)} + \text{CO}_{2(g)} \rightarrow \text{CaCO}_{3(s)}$	2	
[]	$2\text{Na}_{(s)} + \text{Cl}_{2(g)} \rightarrow 2\text{NaCl}_{(s)}$		Reacción de neutralización
[]	$\text{AgNO}_{3(ac)} + \text{NaCl}_{(ac)} \rightarrow \text{AgCl}_{(s)} + \text{NaNO}_{3(ac)}$	3	
[]	$\text{Mg}_{(s)} + 2\text{HCl}_{(ac)} \rightarrow \text{MgCl}_{2(ac)} + \text{H}_{2(g)}$		Reacción de desplazamiento doble o sustitución doble
[]	$\text{CH}_4 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{CO}_2$		
[]	$2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$		
[]	$\text{N}_{2(g)} + 2\text{O}_{2(g)} \rightarrow 2\text{NO}_{2(g)}$		
[]	$(\text{NH}_4)_2\text{Cr}_2\text{O}_{7(s)} \rightarrow \text{Cr}_2\text{O}_{3(s)} + \text{N}_{2(g)} + \text{H}_2\text{O}_{(l)}$	4	
[]	$\text{SF}_{4(g)} + \text{F}_{2(g)} \rightarrow \text{SF}_{6(g)}$		Reacción de desplazamiento simple o sustitución simple
[]	$2\text{AgBr}_{(s)} \rightarrow 2\text{Ag}_{(s)} + \text{Br}_{2(l)}$	5	
[]	$\text{O}_{2(g)} + 2\text{SO}_{2(g)} \rightarrow 2\text{SO}_{3(g)}$		Reacción de combinación
[]	$\text{Ca}(\text{OH})_{2(aq)} + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{2(aq)} + 2\text{H}_2\text{O}_{(l)}$	6	
[]	$2\text{H}_2\text{O}_{(l)} \rightarrow 2\text{H}_{2(g)} + \text{O}_{2(g)}$		
[]	$2\text{HgO}_{(s)} \rightarrow 2\text{Hg}_{(l)} + \text{O}_{2(g)}$		
[]	$\text{HCl}_{(ac)} + \text{NaOH}_{(ac)} \rightarrow \text{NaCl}_{(ac)} + \text{H}_2\text{O}_{(l)}$		