

Limiting Reactant/Reagent Worksheet 2

All of the questions on this worksheet involve the following reaction: When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

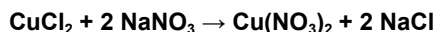
- 1) Write the balanced equation for the reaction given above:
- 2) If 15.0 grams of copper (II) chloride react with 20.0 grams of sodium nitrate, how much sodium chloride can be formed?
- 3) What is the limiting reagent for the reaction in #2? _____
- 4) How much of the nonlimiting reagent is left over in this reaction?
- 5) If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

ANSWERS - Limiting Reagent Worksheet

Limiting Reagent Worksheet

All of the questions on this worksheet involve the following reaction: When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

- 1) Write the balanced equation for the reaction given above:



- 2) If 15.0 grams of copper (II) chloride react with 20.0 grams of sodium nitrate, how much sodium chloride can be formed?

To solve this problem determine how much sodium chloride can be made from each of the reagents by themselves. When you work out how much sodium chloride can be made with 15.0 grams of copper (II) chloride, you find that 13.0 grams will be formed. When starting with 20.0 grams of sodium nitrate, 13.6 grams will be formed. Since 13.0 grams is the smaller number, that's our answer. Please note that rounding differences may cause your answers to be slightly different, so if they are, don't panic.

- 3) What is the limiting reagent for the reaction in #2? **copper (II) chloride**

- 4) How much of the non limiting reagent is left over in this reaction?

Take your 15 grams of limiting Copper (II) Chloride and turn it into sodium nitrate. This will be the amount of sodium nitrate needed. Subtract this number from your starting amount of sodium nitrate of 20 grams.

- 5) If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$11.3/13.0 \times 100\% = 86.9\%$$