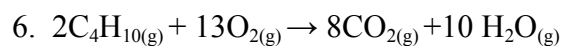
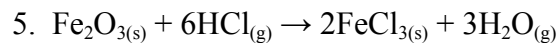
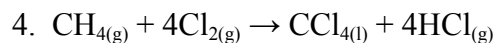
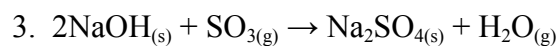
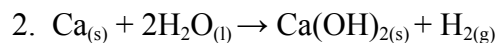
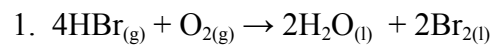


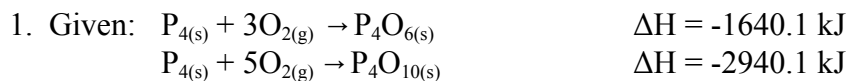
Name _____ Date _____ Hour _____

Thermochemistry- ΔH

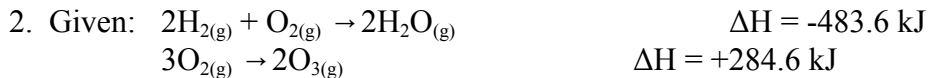
Determine ΔH for the following reactions. State whether they are exothermic or endothermic.



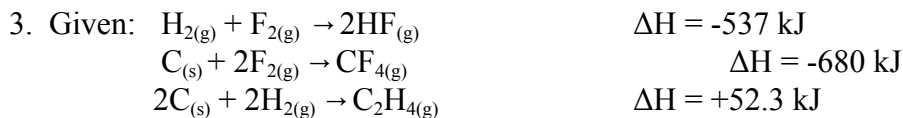
Hess's Law



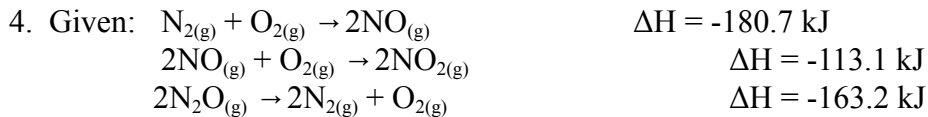
Calculate the enthalpy change for the reaction: $\text{P}_4\text{O}_{6(\text{s})} + 2\text{O}_{2(\text{g})} \rightarrow \text{P}_4\text{O}_{10(\text{s})}$



Calculate the enthalpy change for the reaction: $3\text{H}_{2(\text{g})} + \text{O}_{3(\text{g})} \rightarrow 3\text{H}_2\text{O}_{(\text{g})}$



Calculate the enthalpy change for the reaction: $\text{C}_2\text{H}_{4(\text{g})} + 6\text{F}_{2(\text{g})} \rightarrow 2 \text{CF}_{4(\text{g})} + 4\text{HF}_{(\text{g})}$



Calculate the enthalpy change for the reaction: $\text{N}_2\text{O}_{(\text{g})} + \text{NO}_{2(\text{g})} \rightarrow 3\text{NO}_{(\text{g})}$