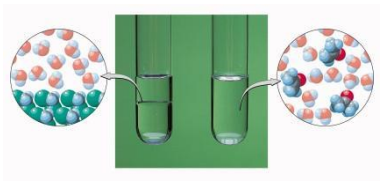
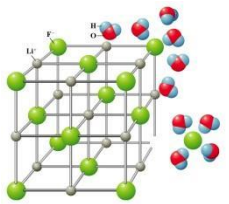


Unit 10:
Properties of Solutions

10-1 through 10-3



LT 10 I can describe how the intermolecular interactions and the mathematical relationships between solute and solvent can explain the physical properties of aqueous solutions.			
Subtarget	*Activity/Lab/Demo*	*Worksheet/Resources*	*Bookwork*
10.1 I can express the energy changes in the solution process			
10.2 I can distinguish between unsaturated, saturated and supersaturated solutions	👩‍🔬Activity-Supersaturated Solution		🍽Supplement: 1,2
10.3 I can explain the factors that affect solubility	👩‍🔬Lab-Solubility of a Salt	✍WS: Solubility Curves ✍WS: Solubility Graph	🍽Read & Notes 10-2 🍽Book: 21,23
10.4 I can calculate the concentration of a solution in molarity, mass percent, mole fraction and parts per million/billion	👩‍🔬Lab-Beer's Law	✍WS: Solution Concentration	🍽Read & Notes 10-1 🍽Book: 1,3,5,7,9abc,11, 13,15,17ab *skip molality*
10.5 I can explain why the physical properties of solutions are dependent on the concentration of the solute and the strengths of all interactions among the particles of the solutes and solvent.	👩‍🔬Activity-Ice Cream		🍽Read & Notes 10-3 (NO MATH) 🍽Supplement: 3,4

Workbook ($3): Lesson 7: Ex 1-3, Prob 4a,5,9,10a
Conceptual Problems ($3): 69-81
No challenge problems

These are practice for the extra credit but NOT required for chem menu credit
🍽Book: 39a,41,43,47,53,59