Lab: Experimental Determination of Calcium Carbonate in Tums

Introduction

- Experiment objective
- Reaction of Tums (active ingredient) with stomach acid:

$$\circ \quad 2 \text{ HCl}_{\text{(aq)}} + \text{ CaCO}_{3 \text{ (s)}} \rightarrow \text{H}_2\text{O}_{\text{(I)}} + \text{ CO}_{2 \text{ (g)}} + \text{ CaCl}_{2 \text{ (aq)}}$$

- Titration technique
 - Endpoint: moles of titrant = moles of analyte (titrant NaOH, analyte HCl)
- Acid-Base reaction during titration:

$$\circ$$
 HCl _(aq) + NaOH _(aq) \rightarrow H₂O _(l) + NaCl _(aq)

Procedure

- Written in passive voice, past tense
- Omit Part 1 it was just for practice!

→ Include the averages you took!

• Remember, this is not the time to include specific data.

Results

Part 1 info is not required – it was just for practice!

Describe in complete sentences where these data came from:

Mass of Tablet (g)

Mass of Samples (g)

Concentration of HCI (M)

Volume of HCI added (mL)

Concentration of NaOH (M)

Volume of NaOH added (mL)

Discussion

Contains calculations along with explanations in complete sentences!

- Calculations 2-5: Describing the acid-base reaction.
 - \circ HCl_(aq) + NaOH_(aq) \rightarrow H₂O_(l) + NaCl_(aq)
 - Based on the endpoint of the titration, you determined how many moles of acid were neutralized when you put in a piece of the tablet.
- Calculations 6-8: Describing the reaction of HCl with CaCO₃
 - \circ 2 HCl _(aq) + CaCO_{3 (s)} \rightarrow H₂O _(l) + CO_{2 (g)} + CaCl_{2 (aq)}
 - Based on the reaction above, you determined how many moles of CaCO₃ must be in the piece of tablet. Then you determined the percentage of CaCO₃ and the mass in the full tablet.
- Calculation 9: Calculating percent difference from the label
 - Compared the mass of CaCO₃ from above to the value listed on the label of the Tums bottle

Conclusion

- Re-state objective
- Summarize the outcome
 - $\circ\quad$ Include the final value: amount of CaCO_3 in tablet and percent difference from the Tums label