

1st Third of 6th Grade Science Vocabulary

1. **Scientific Method** - (or inquiry) a process that uses a set of skills to answer questions or to test ideas about the natural world.
2. **Science** - the investigation and exploration of natural events and the new information that results from those investigations.
3. **Observation** - the act of using one or more of your senses to gather information and taking notes of what occurs
4. **Description** - a spoken or written summary of observations
5. **Qualitative description** - description using your senses (sight, smell, etc.)
6. **Quantitative description** - description using number and measurements
7. **Inference** - a logical explanation of an observation that is drawn from prior knowledge and or experience (the result of observation)
8. **Hypothesis** - a possible explanation for an observation that can be tested by scientific investigations
9. **Prediction** - a statement of what will happen next in a sequence of events
10. **Investigation** -
11. **Explanation** - an interpretation of observations
12. **Scientific theory** - an explanation of observations or events that is based on knowledge gained from many observations and investigations
13. **Scientific law** - a rule that describes a pattern in nature
14. **Critical thinking** - comparing what you already know with the information you are given in order to decide whether you agree with it
15. **International System of Measurement (SI)** - internationally accepted system of measurement
16. **Meter** - SI base unit for length
17. **Liter** - SI unit for volume
18. **Gram** - SI base unit for mass
19. **Celsius** - temperature scale in which 0 represents the freezing of water
20. **Kilo-** prefix meaning 1,000
21. **Hecto** - prefix meaning 100
22. **Deka (Deca)** - prefix meaning 10
23. **Deci** - prefix meaning .10
24. **Centi-** prefix meaning .100
25. **Milli** - prefix meaning .1000
26. **Length** - describes how long something is
27. **Distance** - describes how far away or near one object is to another
28. **Volume** - the amount of space matter takes up
29. **Water displacement method** - a way to measure the volume of an irregularly shaped object using water
30. **Meniscus** - the bottom of the curve formed by a liquid in a graduated cylinder
31. **Mass** - the measure of the amount of matter in an object
32. **Meter stick** - device used to measure lengths and distances
33. **Graduated cylinder** - device used to measure the mass of matter
34. **Triple beam balance** - device used to measure the mass of matter
35. **Goggles** - eye protection equipment worn in the lab
36. **Thermometer** - a device used to measure temperature
37. **Microscope** - a device used to help magnify tiny objects that can not be seen with human eye simply to make tiny features of objects easier to view
38. **Matter** - anything that has mass and takes up space
39. **Atom** - a small particle that is the building block of matter
40. **Nucleus (chemical)** - the region in the center of an atom where most of an atom's mass and positive charge is concentrated
41. **Proton** - the positive particle in the nucleus of an atom
42. **Neutron** - a neutral in the nucleus of an atom
43. **Electron** - a negatively charged particle that occupies the space in an atom outside the nucleus
44. **Atomic number** - the number of protons in the nucleus of an atom of an element
45. **Atomic mass** - the number of protons plus the number of neutrons
46. **Element** - substance that consists of only one type of atom
47. **Periodic table** - a chart of the elements arranged into rows and columns according to their physical and chemical properties
48. **Pure substance** - matter with a composition that is always the same

50. **Chemical properties** - the ability of a substance to combine with or change into one or more new substances.
51. **Melting point** - temperature at which a solid changes to a liquid
52. **Boiling point** - temperature at which a liquid changes to a gas
53. **Precipitate** - a solid that sometimes forms when two liquids combine
54. **Ductility** - the ability of a substance to be pulled into thin wires
55. **Malleability** - the ability of a substance to be hammered or rolled into sheets
56. **Conductivity** - able to transfer electricity and thermal energy
57. **Insulator** - any material that keeps (prevents) energy such as electricity, heat, or cold from easily transferring through it
58. **Luster** - the way a mineral reflects or absorbs light at its surface
59. **Density** - the mass per unit of volume of a substance
60. **Solubility** - the ability of one material to dissolve in another
61. **Dissolve** -
62. **Hardness** - the resistance of a mineral to being scratched
63. **Streak** - a powdery residue produced by minerals when rubbed across and unglazed porcelain tile.
64. **Adamantine** - a brilliant luster such as that of a diamond
65. **Vitreous** - luster that can be described as glassy (looks like glass)
66. **Pearly** - luster with the appearance of a pear, play of colors
67. **Metallic** - luster that can be described as shiny like a metal
68. **Mohs Hardness Scale** - a scale developed by German mineralogist Friedrich Mohs to compare the hardness of different minerals
69. **Chemical change** - a change in matter in which the substances that make up the matter change into other substances
70. **Physical change** - a change in matter in which the substance that make up matter does not change the matter's identity.
71. **Law of Conservation of Mass** - law that states the total mass before a chemical reaction (change) is the same as the total mass after the chemical reaction.
72. **Metal** - an element that is generally shiny, is easily pulled into wires or hammered into thin sheets, and is a good conductor of electricity and thermal energy.
73. **Metalloid** - an element that has physical and chemical properties of both metals and nonmetals.
74. **Physical properties** - a characteristic of matter that you can observe or measure without changing the identity of the matter
75. **Halogens** - (HA luh jun) and element in group 17 on the periodic table. (from the Greek hals, means "salt"; and 'gen, means to produce)
76. **Noble gas** - an element in group 18 on the periodic table only reacts with other elements under special lab conditions free elements; not compounds naturally.
77. **Semiconductor** - a substance that conducts electricity at high temperatures but not at low temperatures.
78. **Nonmetal** - an element with no metallic properties (Ex: Carbon, Phosphorus, Chlorine, & Helium)
79. **Compound** - substance containing two atoms when two or more different elements
80. **Molecule** - two or more atoms that are held together by covalent bonds and acts as a unit
81. **Chemical formula** - the combination of chemical symbols that represent the name of a molecule or compound.

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