

Thermochemistry Vocabulary

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Word	Definition
Endothermic	<p>In an endothermic reaction the energy to break bonds is larger than the energy released when bonds are formed. The overall reaction absorbs energy so the products have more energy than the reactants.</p> <p>One example reaction is melting ice. The water can move more when it is melted than it can when frozen so it contains more energy. It feels cold because the ice is stealing energy from your hand to melt.</p>
Exothermic	<p>In an exothermic reaction the energy to break bonds is smaller than the energy released when bonds are formed. The overall reaction releases energy so the products have less energy than the reactants.</p>
Bond energy	<p>The bond energy is the energy it takes to break a bond.</p>
Specific Heat	<p>The specific heat is the energy it takes to raise one gram of a substance by one degree.</p>
Energy transfer	<p>When high motion particles move some of their motions to slower moving particles. For example, heat travels from hot objects (high motion) to cold objects (slow motion).</p>
Energy transformation	<p>An energy transformation happens when energy is converted from one form to another. For example in a power plant the thermal energy from burning fuel is converted to electrical energy</p>
Activation energy	<p>The energy added to break bonds of a reaction and get the reaction started.</p>
<u>Review vocabulary needed for this unit</u>	
Word	Definition
Products	<p>The elements or molecules formed in a reaction. Products are shown on the right side of the arrow in the equation.</p>
Reactants	<p>Elements or molecules that start a reaction. Reactants are shown on the left side of the arrow in the equation.</p>
Coefficients	<p>The large numbers before a molecule in a reaction. These indicate how many molecules or elements of each type are involved in a reaction</p>
subscripts	<p>Small numbers within a chemical formula that indicate how many of each type of element or ion are within that molecule or compound.</p>

