Electronegativity and Polarity Notes

Ionic-

Covalent-

nonpolar covalent-

polar covalent

Electronegativity-

Electronegativity Trends

Electronegativity Differences

	Non-Polar Covalent	Polar Covalent	Ionic
Specific Rule:			

For each pair draw an arrow under the molecule showing its **dipole moment**. Or write “none” if there isn’t one.

H-Cl H-H H-I Br-Br C-O C-F Cl-N

Bond Strength and Bond Energy-

What's a Bond?

Ionic

Covalent

For each of the following pairs of compounds, circle the pair that is more polar

H-S or H-F

O-S or O-F

N-S or N-Cl

C-S or C-Cl

Label the following as Ionic or Covalent using basic rules (M+NM or NM+NM)

CO_2 NaCl H_2O $MgCl_2$ NO_3

What is the difference in electronegativity between each and label them from most polar (1) to least polar(7)

	N-S	Cl-Cl	H-Si	Li-F	P-Cl	S-O	C-H
E dif	____	_____	_____	_____	_____	_____	_____
rank	____	_____	_____	____	_____	_____	_____
I/PC/NPC	____	_____	_____	____	_____	_____	_____

Intramolecular Forces

Intermolecular Forces