

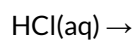
Acids and Bases
Introduction

Definitions

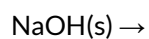
Arrhenius acids and bases

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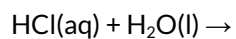
The terms “hydrogen ion” and “hydronium ion” and the symbols $\text{H}^+(\text{aq})$ and $\text{H}_3\text{O}^+(\text{aq})$ are often used interchangeably for the aqueous ion of hydrogen. Hydronium ion and $\text{H}_3\text{O}^+(\text{aq})$ are preferred, but $\text{H}^+(\text{aq})$ is also accepted on the AP Exam.



Bronsted-Lowry acids and bases

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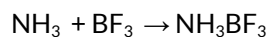
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Lewis acids and bases

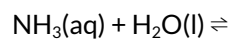
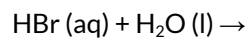
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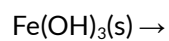
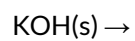


Ionization and Dissociation

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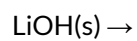
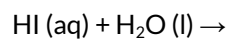
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Strong Acids and Bases

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Strong Acids

Hydrochloric Acid

Nitric Acid

Hydrobromic Acid

Perchloric Acid

Hydroiodic Acid

Sulfuric Acid

Strong Bases

Lithium hydroxide

Sodium hydroxide

Potassium hydroxide

Rubidium hydroxide

Calcium Hydroxide

Strontium hydroxide

Barium hydroxide

Weak Acids and Bases

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