

Ideal gases

Gases whose particles bounce off each other (and the walls of their containers) with no loss in energy, have no intermolecular forces of attraction, do not react or interact with each other chemically and are essentially “point particles” (no size). THERE IS NO SUCH THING AS AN IDEAL GAS.

Never-the-less, most gases almost behave as ideal gases most of the time. An ideal gas would perfectly obey the gas laws. Most gases very “closely” obey the gas laws.