

Balancing Equations and Types of Reactions Review

Types of Reactions

- 1) $\underline{\hspace{1cm}} \text{Ag}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{Ag} + \underline{\hspace{1cm}} \text{O}_2$ $\underline{\hspace{4cm}}$
- 2) $\underline{\hspace{1cm}} \text{KClO}_3 \rightarrow \underline{\hspace{1cm}} \text{KCl} + \underline{\hspace{1cm}} \text{O}_2$ $\underline{\hspace{4cm}}$
- 3) $\underline{\hspace{1cm}} \text{H}_2 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{O}$ $\underline{\hspace{4cm}}$
- 4) $\underline{\hspace{1cm}} \text{K} + \underline{\hspace{1cm}} \text{MgBr}_2 \rightarrow \underline{\hspace{1cm}} \text{KBr} + \underline{\hspace{1cm}} \text{Mg}$ $\underline{\hspace{4cm}}$
- 5) $\underline{\hspace{1cm}} \text{AlBr}_3 + \underline{\hspace{1cm}} \text{K}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{KBr} + \underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3$ $\underline{\hspace{4cm}}$
- 6) $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$ $\underline{\hspace{4cm}}$
- 7) $\underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{F}_2 \rightarrow \underline{\hspace{1cm}} \text{NaF} + \underline{\hspace{1cm}} \text{Cl}_2$ $\underline{\hspace{4cm}}$
- 8) $\underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{NH}_3$ $\underline{\hspace{4cm}}$