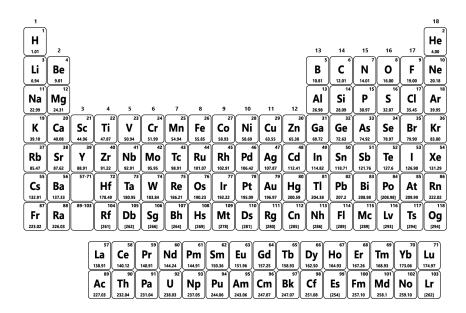
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Physical Science Unit 5 Study Guide

Ch	apter 5:
Kno	ow all of the vocabulary from Chapter 5.
1.	How did Mendeleev arrange his periodic table?
2.	How is the modern periodic table arranged?
3.	What is a group?
4.	What is a period?
5.	What is atomic mass and what is the unit?
6.	Where are the metals on the periodic table? Where are the nonmetals?
7.	Which two groups contain the most reactive elements?
8.	Which one group contains the least reactive elements?

What is the reactivity of an element based on?

List the number of valence electrons in each group on the periodic table below. (Only label groups 1, 2, and 13 – 18.)



- Within group 1, what is the most reactive element?
- Within group 17, what is the most reactive element?
- Why are noble gases so stable?
- 14. Name the following groups:
 - a. group 1

a stair-step elements

_{b.} group 2

group 17

c. groups 3-12

f. group 18

Label each element as a metal, nonmetal, or metalloid.							
	a.	hydrogen	d.	iron			
	b.	calcium	е.	silicon			
	c.	bromine	f.	neon			
Ch	apte	r 6:					
(no	ow a	ll of the vocabulary words from Chapter	· 6.				
16.	Why	y do certain atoms lose or gain electrons	ŞŞ				
17.	Whi	ch elements have the most stable elect	ron	configurations?			
18.	Who	at is the difference between an ionic bo	nd	and a covalent bond?			
19.	What are three properties of ionic compounds?						
20.		v can you tell from a formula if a compo ecules?	oun	d is made of ions or			
21.	Hov	v are the charges of transition metals dif	fere	ent from other elements?			
22.	Nar	me the following compound: $MgCl_2$.					

23.	What does the roman numeral mean in the name iron (III) oxide?				
24.	When aluminum and fluorine bond, what is the formula?				
25.	What is the formula for diphosphorus trioxide	÷\$			
26.	What is a metallic bond?				
27.	Why are metals good conductors and malleable?				
28.	Draw the electron dot diagram for the following elements.				
	a. H	c. P			
	b. Br	d. Ba			
29.	What is the name of SO ₃ ?				
30.	Draw an example of a single bond, double	bond, and triple bond.			