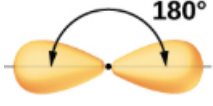

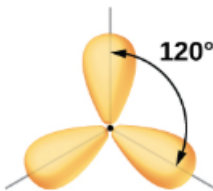

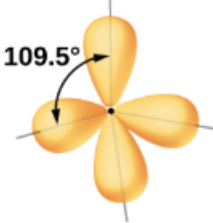

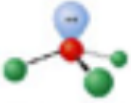
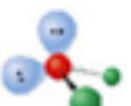
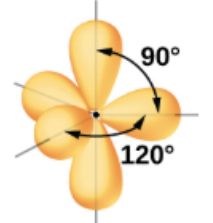



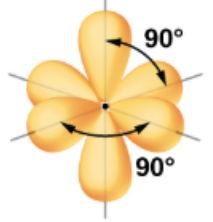


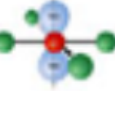
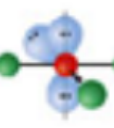



VSEPR Theory Geometry Chart

Number of Electron Dense Areas	Electron Domain Geometry Name	Molecular Geometry Name				
		0 Lone Pairs	1 Lone Pair	2 Lone Pairs	3 Lone Pairs	4 Lone Pairs
	Linear	 Linear				
	Trigonal Planar	 Trigonal Planar				
	Tetrahedral	 Tetrahedral	 Trigonal Pyramidal	 Bent		
	Trigonal Bipyramidal	 Trigonal Bipyramidal	 Seesaw	 T-Shaped		
	Octahedral	 Octahedral	 Square Pyramidal	 Square Planar	 T-Shaped	 Linear

Bond Polarity is determined by the electronegativity difference between two atoms. The greater the difference in electronegativity, the more polar the bond. The smaller the difference in electronegativity, the less polar the bond.

Molecular Polarity is determined by the arrangement of the electron domains around the entire molecule. If the domains cancel each other out (very symmetrical), the molecule is non-polar. If the domains do not cancel each other out (asymmetrical), the molecule is polar.