Beanium

pg. 113:
#19- Three isotopes of oxygen are oxygen-16, oxygen-17, and oxygen -18. Write the symbol for each, including the atomic number and mass number.
#20- three isotopes of chromium are chromium -50, chromium-52, and chromium-53. How many neutrons are in each isotope, given that chromium has an atomic number of 24?

Pg. 116:
#21- Boron has two isotopes: boron-10 and boron-11. Which is more abundant, given that the atomic mass of boron is 10.81?
#22- There are three isotopes of silicon; they have mass numbers of 28,29, and 30. The atomic mass of silicon is 28.086 amu. Comment of the relative abundance of these three isotopes.

Pg. 117:
#23- The element copper has naturally occurring isotopes with mass numbers of 63 and 65. The relative abundance and atomic masses are 69.25% for mass=62.93 amu, and 30.8% for mass =64.93 amu. Calculate the average atomic mass of copper.