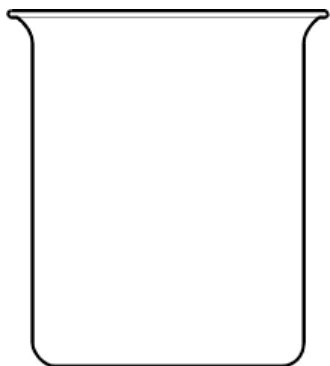


Titrations WS

1. What is the endpoint of a titration? (Consider phenolphthalein and bromothymol blue)
2. What is the equivalence point of a titration?
3. It takes 53.5 mL of a 0.45 M NaOH solution to neutralize 185 mL of an HCl solution.
 - a. What is the concentration of the HCl solution?

- b. At the ***endpoint***, sketch the species that are in the beaker.



4. You are titrating an acid into a base to determine the concentration of the base. The endpoint of the neutralization is reached but the stopcock on the buret sticks slightly and allows a few more drops of acid to fall into the solution. How will this affect your calculations for the concentration of the base?

